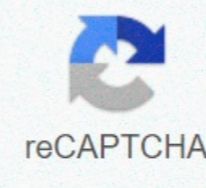




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Periodic classification of elements class 10 worksheet with answers

The elemental classification of the Periodic Element of the Mendeleef Periodic Law depends on - A. Atomic Weight B. The number of neutron D. No atoms, the heaviest among the following are A. U B. Ra C. Pb D. Hg, which pairs are both members from the same period table? The diagonal relationship is expressed by - A element of the first interval and the composition of the second interval B, the composition of the second interval, and the composition of the third period C. Element of the third period, and the composition of the fourth period D. All the elements with atomic number 36 belongs to ... the block in the periodic table A. B. C. F D. which the following shows the diagonal relationship? A.B and Si B.B and Al C.B and Ga D.B and C. A. Na, K, Rb B. Mg, S, As C. Cl, Br, I D. P, S, are the last members in each moment of the table is _____. A. Noble Gas Composition B. Transition Element C Halogen D. Alkaline metals, one of the following combinations, represent the metal elements A. 2, 8, 7, 2, 8, 8 C. 2, 8, 4 D. 2, 8, 2, 2, 2, 2, long form of periodic table on _____ are based on the electronic configuration of atoms B of atomic weight of the physical properties of the element d. Electronegativity Key Answer: Description Fact: U> Ra& Ra. Leaving an atom designated uranium atom is the heaviest. Description: Na – Cl. Both are of the third period Description: Elements of diagonal relationship second moment and three description: Kr There is an atom number 36, which is a noble gas, and all the noble gases are of p-block Description: According to Doberner's rule, the atomic mass of the central element is almost the arithmetic mean of the atomic mass of two other elements. : atoms easily loosen electrons. The element classification of worksheet-4 atomic elements is the smallest.C? If the Valenine's shell of the atom of the element has 7 electrons, the composition belongs. of _____ As we move along the period of _____ B. Increase B. The first increase C. The first increase is reduced D. Remains constant at any time, the valency of the element associated with oxygen _____ is of increased one by one from the vlia of one by one by one. One by one from IA is IVA and then increased from VA to VIIA one by one, which shows the highest non-metallic character.B? The chemical behavior of atoms is determined by _____ as of the number of atoms B of mass B of binding energy d. the number of isotopes, which follows the noble gas composition? A. Na B. Fe C. Li D. The lightest metal is _____. A. Li B. Mg C. Ca D. Na's following pair, one with an example of metalloid elements in the is _____ periodic table belonging to sodium and potassium B. fluorine and chlorine c. of calcium and magnesium deboron and silicon, the number of elements in the longest period of the is _____ table of A. 2 B. 8 C. 18 D. 32 Answer Key: The size of the atom is reduced from left to right. Metal oxide is fundamental in nature and non-metallic is acidic in nature. Description: Description: Description Facts: The chemical behavior of atoms depends on the number of Valentine electrons. Description: Description Facts: Li, since it is the smallest. So they are metalloids to the question of the board. CBSE 1. (a) What is metal? Write two examples. (b) listed below are some elements of the first group Li, Na, K (their atomic numbers are 3,11,19, respectively, and are of the 2, 3 and 4, respectively). These sort of metal characters in the lower order shown by them 2.What is amposporic oxide? Select the following oxide amps from the following oxide groups. :- Na2O, ZnO, Al2O3, CO2, H2O 3. of the element 'X' in the periodic table, (ii) write the formula of the compound that occurs when the 'X' reacts/combines with another element 'Y' (atomic number 8) (iii) what would be natural (acidic or basic) of the resulting compound? Adjust your answer well [2015] Ans: Electronic configuration: 2,8 8 , 2 1/2) 'X' is in group 2 and period 4 of the periodic table 1/2, 1/2 ii) XY 1/2/iii) basis because X is a metal and metal oxide of nature (Y, atomic number = 8 , oxygen is basic) 1/2, 1/2, 1/2, 1/2 The four elements P, Q, R and S are the third period of the modern-day grid and contain electrons 1, 3, 5 and 7, respectively, in their outermost shells. Write the electronic configuration of Q and R and define their Valentine. Write the molecular formula of the compounds that occur when P and S include Ans: Electronic configuration of Q: 2, 8, 3 Valency of Q: 3 Electronic configurations of R: 2, 8, 5 Valency of R: 8 – 5 = 3 Electronic configurations, Number P: 2, 8 , 1 Electronic configuration of S: 2 , 8 , 7Q Two elements 'P' and 'Q' are in the same period of modern periodic table and belong to group 1 and group-2 respectively. Compare the following characteristics in the (a grid model), the number of electrons in their atoms, (b) the size of their atoms, (c) their metal character(d) their tendency to lose electrons. (e) formulated of their oxides (f) Formula of their chloride [2015]Ans: P Q(a) No. of electrons in atoms 3 4 11 12 19 20 (b) The size of the atom is bigger. Smaller metal characters (c) fewer metal characters (d) tend to lose fewer electrons(e) the formula of the oxide P2O QO (f) formula of chloride. PCI QCl2 Q.Using an example of atomic composition no. 16 explains how the electronic configuration of the atom of the element relates to the position in the modern periodic table and how the fluency of the element is calculated on the basis of the atom [2015]. Since the number of electron valence is 6 group, the number will be 10 + 6 = 16 Valenc. Ans: This is because there are no atoms, all the elements are unique, but the atomic mass may be the same as isotopes Q. A quiz contest is held in schools for chemistry students. (b) What elements of N, F, P and Ar show similarities to this element, (c) We use the compounds of this element in our diet every day. That's what Ans: (a) atomic number =17 (b) F- because it is the same group as the composition (C) (c)NaCl-Common Salt (d)Ca(HCO3)2, creating a temporary active energy of the water. Water

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