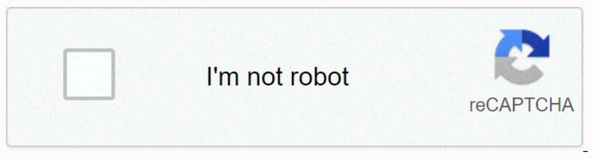

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Combustion definicion rae

The concept of oxygen therapy is not part of the Dictionary of the Royal Spanish Academy (RAE). This concept develops with the combination of two terms: oxygen and therapy. Oxygen is a gaseous chemical element found in the atmosphere and in all living things. It is indispensable, for example, for breathing and for the development of combustion. Therapy, meanwhile, is a treatment that is carried out in order to treat a disease, a condition of another kind or a psychological problem. With these thoughts in mind, we can move forward with the definition of oxygen therapy. This is called therapy, which uses oxygen. It is the supply of this substance in a concentration higher than the concentration in the air of the environment, with the intention of increasing the concentration of oxygen present in the blood and thus preventing injury caused by hypoxia (lack of oxygen in the body). Before appealing to oxygen therapy, the patient is usually exposed to arterial gasometry or pulsioxymetry. In this way, you can know the concentration of oxygen in the blood. If a person can not maintain the correct level of oxygen, he has respiratory impairment. Oxygen therapy can be carried out with masks or cannulas attached to an oxygen tube or hyperbaric chamber. In this way, the body receives gas with a flow indicated by a doctor or nurse. It should be noted that oxygen therapy is also used in the field of aesthetics. In this case, the treatment aims to improve the appearance of the skin, since oxygen moisturizes and also increases collagen production. In this framework, we can talk about oxygen therapy in the face, which is one of the most modern trends in cosmetic procedures to improve the appearance and shine of the skin of the face. As is often the case in this area, when a music or movie star publicly declares that he has resorted to a new treatment, his popularity rises across the sky as everyone wants to try, and it has been the same with facial oxygen therapy thanks to Cameron Diaz and Madonna. The advantages of oxygen treatment in the face are several. First of all, we should mention that the hydration of the skin of the face is given at a deep level, and therefore improving its structure and tension takes a lot. On the other hand, this treatment also helps to prevent skin peeling and redness. It is also worth noting the antioxidant force provided by oxygen teritodation of the face. Free radicals are among the most common causes of premature skin aging, and thanks to antioxidants we can fight them to minimize their effect and achieve two hard-to-resist benefits: the appearance of wrinkles and prevent the appearance of new ones. To treat problems such as rosacea and acne, oxygen therapy benefits us by stimulating blood circulation. Thanks to this action, the skin of the face receives a greater amount of nutrients, collagen and vitamins and recovers more often. Collagen is a key protein for the skin to stay firm, which is why this treatment is so suitable for the face, the part of the body that first shows signs of aging. While wrinkles are a natural element of growth, there is a less common problem than facial oxygen therapy can also help solve: spots on the skin. In this group, we need to include some markers and scars that can benefit from regenerative action. Consultation possible thanks to the commitment to the culture of combustion Del lat. combustio, -Nis. 1. f. The action and effect of burning or burning. 2. f. Biol. Oxidation of food in living things. 3. f. Chim. Chemical reaction between oxygen and oxidizing material, accompanied by energy separation and usually manifested by a light bulb or flame. nuclear combustion 1. F. A set of nuclear reactions that lead to the continuous production of large amounts of energy. combustion chamber internal combustion engine Real Academia Española © All rights reserved Discover a new platform of language resources RAE cotillón Thursday, December 31, 2020 Panhyspanic Dictionary Legal Spanish Amb. Oxidation of fuels, regardless of the use of heat or electrical or mechanical energy produced by this process, and any other directly connected activity, including washing of waste gases. Ley 1/2005, de 9-III, por la que se regula el régimen del comercio de derechos de emisión de gases de efecto invernadero, art. 2.combustible sólido nacional 氧化剂 1.325 millones de hablantes comburente 570 millones de hablantes Oxidizing 510 millones de hablantes आक्सीकारक 380 millones de hablantes 280 مؤكسد millones de hablantes окислитель 278 millones de hablantes oxidante 270 millones de hablantes জারক 260 millones de hablantes comburant 220 millones de hablantes pengoksida 190 millones de hablantes Oxidations 180 millones de hablantes 酸化剤 130 millones de hablantes 산화 제 85 millones de hablantes oxidizer 85 millones de hablantes oxidizer 80 millones de hablantes ஆக்கிஜனாக்கி 75 millones de hablantes oxidizer 75 millones de hablantes yükseltgen 70 millones de hablantes ossidante 65 millones de hablantes utleniacz 50 millones de hablantes окислювач 40 millones de hablantes oxidant 30 millones de hablantes οξειδωτής 15 speakers oksideermiddel 14 million speakers oxidationsmedel 10 million speakers oksidant 5 million speakers 点the 1.325 million speakers ignition 570 million speakers ignition 510 million speakers ইগ্নিশন 380 million speakers 280 اووال million speakers 2 2 2 00 78 million speakers igni-about 270 million speakers ইগনিশন 260 million speakers on all 220 million speakers pencucuhan 190 million speakers Zandung 180 million speakers 点and 130 million speeches 점- 85 million speakers conk 85 million speakers b. phận lửa 80 million speakers பற்றவைப்பு 75 million speakers प्रज्वलन 75 million speakers 70 million speakers have access to 65 million speakers zapon 50 million speakers 40 million speakers squeeze 3 15 million speakers ontsteking 14 million speakers ant-ndning 10 million speakers tenning 5 million speakers Burning copper (from Latin combustion can generally be understood as any chemical reaction , of a relatively rapid, remarkably exothermic nature, which develops in the gaseous or heterogeneous phase (liquid gas, solid gas) with or without manifestations of flames or visible radiation. [2] From the point of view of classical theory, combustion refers to the rapid oxidation reactions of materials called fuels, which consist mainly of carbon (C) and hydrogen (H) and in some cases sulphur (S), in the presence of oxygen, called an oxidizing agent, and with a large thermal separation. From a functional point of view, combustion is a set of chemical-physical processes through which part of the internal energy of the fuel (chemical energy) is manifested on the outside in the form of heat, which is used inside the oven or boiler. [3] From now on, everything is about classical theory. In fact, instead of pure oxygen, the reaction occurs with the presence of air, which is usually considered the volume of composition to simplify calculations; 21% oxygen and 79% nitrogen. The reactions that occur are as follows: C + O 2 ⟶ CO 2 + 33875 kJ / kg C ' displaystyle ' ce 'C + O2 -> CO2 + 33875 kJ/kg'; H 2 + 1 2 O 2 ⟶ H 2 O + 143 330 kJ / kg O 2 ' displaystyle ' ce 'H2 + 1/2 O2 -> H2O + 143 330 kJ/kg'; O2 S + O 2 O 2 ⟶ OS 2 + 8958 kJ / kg S ' displaystyle ' ce 'S + O2 -> SO2 + 8958 kJ/kg'; S- Combustion Phase Play multimedia content Experiment that demonstrates a large amount of energy due to the combustion of ethanol. 2016 Combustion reactions are actually much more complex than they may seem, especially due to the huge speed at which different phases occur. Even the simplest flame is the result of many almost simultaneous chemical reactions, the study of which requires solving aerodynamic problems, heat conduction and molecular diffusion. [4] Classical theory simplifies this whole process by taking more care of the final result than the dynamics of the process. The three components are not found as pure ingredients, but are part of a compound commonly known as hydrocarbon. The combustion process is carried out in three stages: In the first stage, a prerection occurs, in which hydrocarbons decompose to react with oxygen and form unstable compounds, which are called radicals. The second stage is oxidation, in which most of the heat is released. In the third, oxidation is completed and stable products are formed, which will be components of flue gases. In the first stage, the formed radicals are very active and extremely unstable, so there are chain reactions in which they develop and disappear in a balanced way. When radicals form at a rate higher than they react later, their buildup causes a massive and violent reaction with oxygen known as an explosion. A shock wave caused by a sudden release of energy can reach transmission speeds of more than 2500 m/s and is usually accompanied by detonation. If the rate of propagation is lower than the rate of sound propagation, there will be no explosion, and the sudden reaction is known as deflagration. Combustion types Three types of combustion can be distinguished: Complete or perfect combustion: If the reactions indicated are completely shifted to the right, that is, the components are completely oxidized, forming carbon dioxide (CO2), liquid water (H2O) and possibly sulphur dioxide (SO2), regardless of the amount of air used in the reaction. This means that the oxygen present in the air was at least enough to completely oxidize the components. Stonic or neutral combustion: is a complete combustion in which the exact amount of air obtained by quantitative relationships of the molecules involved in each reaction has been used. It is actually an ideal combustion, which can only be achieved in the laboratory. Incomplete combustion: It is one whose combustion gases contain partially oxidized compounds called unburned, such as: carbon monoxide (CO), carbon particles, Etc. Excess air Combustion reaction C can be written as follows: C + O 2 ⟶ CO 2 'displaystyle'ce 'C + O2 -> CO2' - Quantity by weight and volume, under normal conditions at 0oC and 101.3 kPa pressure, which interferes with each substance in response: 12 grams of carbon need 22.4 liters of oxygen to obtain 22.4 liters of carbon dioxide. If it happens exactly like this, it's what's called a stochyometric reaction. From the estechiometric amount of oxygen and taking into account that the air contains it in 21% of its volume, we can calculate the minimum air needed to provide estechiometric oxygen. For a fuel containing a percentage by mass of carbon, hydrogen and sulphur, the oxygen required to oxidize each element shall be calculated and the sum of these quantities shall be the minimum quantity required to fully burn one unit. [6] In practice, as has already been said, the provision of an combustion process with this minimum amount of air is not achieved by full combustion. It should not be forgotten that in the industrial process both fuel and air are in motion, which makes the mixture of fuel and fuel homogeneous. If added to it; reaction speed, short stay of the mixture in the chamber and temperature fluctuations, as a result, a fraction of the fuel escapes through the chimney without burning, or at least not completely. When carbon particles do not find enough air to burn, the reaction that occurs is: C + 1 2 OR 2 ⟶ CO + 10204 kJ / kg C ' displaystyle ' c + 1/2 O2 -> CO + 10204 kJ/kg; This means that 23 671 kJ is lost for each kilogram of C that passes to CO, except for the risk of carbon monoxide formation. Therefore, it is necessary to provide excess air (n),[7], that is, a greater amount of symmetrically necessary that all fuel particles find enough oxygen to fully oxidize. It could be concluded that securing a large amount of air ensures complete combustion and therefore is a good strategy. However, all the added air is not necessary, it enters and leaves the combustion chamber without doing anything other than heating, which means loss of heat and therefore a decrease in performance. An effective thing, therefore, will add excess air only to achieve complete oxidation. Flue gases Fumes or flue gases are formed by gases resulting from reactions; carbon dioxide, water vapour and sulphur oxide, if any, as well as nitrogen corresponding to the volume of air used. oxygen and nitrogen are found from unused air, nitrogen oxides and sometimes other gases that may be part of the air supplied. Its calculation is carried out in the same way as indicated for oxygen, from combustion reactions and counting excess air. The sum of the volumes obtained for each gas plus the nitrogen contained in the minimum air, all multiplied by the coefficient (n) of excess air, results in the volume of wet gases. If the generated water is not counted in total, the volume of dry gases is obtained. [9] The ratio of CO2 to the total volume of dry gas is the concentration of this dry-based component, which is the maximum that can be measured in normal flue gas analyses. Getting in the analysis of a lower value means: That the excess air was used larger than necessary, because even if the amount of CO2 is as planned, the concentration decreases because the volume of gases in which it is diluted increases. In this case, O2 will also appear in the analysis, the more excess air. Or, some of the carbon is not oxidized, usually by an air defect, in which case the amount of carbon monoxide (CO) appears in the analysis, H2 can appear and with a large amount of air defect does not detect carbon particles (sozo). Excess carbon monoxide air may also occur, especially in liquid fuels, due to poor fuel spraying and a fuel and fuel mixing defect. Calorific value This part is calorific extract [edit] Calorific value is the amount of energy per unit mass or unit volume of mass that can be released when a chemical oxidation reaction occurs. Calorific value expresses energy that can release the chemical bond between fuel and fuel and equals the energy that held atoms together in fuel molecules (combine energy), minus the energy used in the formation of new molecules in materials (usually gases) created in the combustion theory of flames Flames are combustion reactions in motion , at speeds lower than sound. A flame can be defined as an combustion reaction that spreads through space at a lower rate than sound. The concept of flame means movement, and therefore the anterior front reaction called flame front. The shape of a flame or burning fuel depends on the technical medium that prepares the fuel, fuel, mixture of both and provides activation energy,[10] which is the burner. In order to incinerate, to reach the ignition temperature,[11] much higher than at the flash point, which is the one where the fuel is able to start combustion, but if the activation energy is removed, the flame goes out. Of which, all combustion reactions at their different stages are held in this gaseous medium, which is flame. Once ignited, if sufficient fuel and fuel are available at the same rate as the flame, the flame shall stabilise and persist even if the initial activation energy is removed. The rate at which the flame spreads depends on each fuel, its greater or lower degree of purity and the excess air with which it is ching. The temperature of the flame depends on the calorific value, excess air and the type of domestic or combustion chamber. [12] It is known as the theoretical temperature of an adiabatic flame at the highest temperature that can be obtained from fuel that is obtained when there is no excess air. [13] In terms of colour, fuel matters. In general, it should be bright and luminous, without smoldering products. When there is a lack of air, the flame darkens and moisturizes due to hot carbon particles. See also Deflagration Detonation Phlogist Theory Reduction-Oxidation Internal Combustion Engine Irreversible Reaction Breathing Complete Combustion Reference , Combustion, Spanish Language Dictionary (twenty-second edition), Royal Spanish Academy, 2014. Giuliano Salvi. Combustion, p. 4. Camps. Technical manual on the use of liquid fuels in industry. Giuliano Salvi. Combustion (theory and application), p. 158. • Molar volume: the volume occupying the mole, in this case 32 grams of oxygen, or 44 grams per carbon dioxide. In many texts, this minimum required air value is called fuel comious power. • Ratio of excess air (n) to the relationship between the air actually used and the required estechiometrically or minimum n'a r e and l / A m'displaystyle n'A_'real'/A_'m' and excess air 'n' 1 ∗ 100'displaystyle'n-1*100' v %. Ricardo García San Jose. Pollutant emissions. This concept is also known as wet or dry smoke power as or unformed water is considered. • In burners usually obtained by an electrical spark between two electrodes, in individual gas boilers it is obtained by pilot flame, hot wire, etc. In the most common burners of heating boilers, the flame can reach around 1700oC. This temperature can be exceeded by preheating fuel and fuel. John R. Howell. Thermodynamic principles for engineers, p. 496. Bibliography of Giuliano Salvi. Combustion (theory and application). Dossat S. A. ISBN 8423704254. Camps. Technical manual on the use of liquid fuels in industry. Catalan gas and electricity. Energy saving in industry (rational use of fuels). Index. ISBN 8470872257. John R. Howell & Richard O. Buckius. Thermodynamic principles for engineers. McGraw Hill. ISBN 9684225717. Template:HS External Links Wiktionary has definitions and other information about combustion. Data: Q133235 Multimedia: Combustion Obtained from « «

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