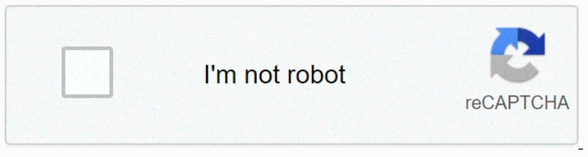


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# Metal reactivity series class 10

A series of metal reactivity, also known as a series of activities, refers to the order of metals in their order of reactivity. Data provided by the reactivity series can be used to predict whether metals can charge another in a single shift reaction. It can also be used to obtain information on the spread of metals towards water and acid. A series chart of regular metal activation is provided below. Metals tend to easily lose electrons and form decorations. Most of them react with atmospheric oxygen to form metal oxide. However, different metals have different reactive towards oxygen (inactive metals such as gold and platinum do not easily form oxide when exposed to air). An important feature of Metal at the top of the series of reactivity is that agents reduce strong ones because they are easily oxidized. These metals are tarnish/erring easily. Reducing metal capabilities grows weaker while traveling down the series. Electro positivity of the elements also reduces while moving down a series of metal activations. All metals found on hydrogen in a series of activities free up H2 gas when reacting with HCl melted or melting H2SO4. Metals placed higher on a series of reactivity have the ability to keep metals placed lower than their salt solutions. Higher ranking metals need more energy for their isolation from ores and other compounds. Another important feature of the series of activities is that while traveling down the series, the ability to donate metal electrons reduces. Long Tabular Form Series Metal Reactivation Reactivation is marked below (in descuming order) along with the corresponding ions. Note that metals in Red react with cold water, those in orange can not react with cold water but can react with acid, and those in blue react only with some powerful oxidizing acid. Caesium Cs+ Francium Fr+ Rubidium Rb+ Potassium Series K+ Sodium Na+ Lithium Lithium Li+ Barium Ba2+ Radium Ra2+ St Rontium Sr2+ Calcium Ca2+ Magnesium Mg2+ Be2+ Aluminum Al3+ Titanium Ti4+ Manganese Mn2+ Zinc Zn2+ Ch Cr3+ Iron Fe3+ Cadmium Cd2+ Cobalt Co2+ Nickel Ni2+ Tin Sn2+ Lead Pb2+ Hydrogen H+ (Non-Metals, References for Comparison) Antimoni Sb3+ Bismuth Bi3+ Copper Cu2+ Tungsten W3+ Mercury Hg2+ Silver Ag + Platinum Pt4+ Gold Au3+ Despite being non-metallic, hydrogen is often included in a series of reactivity because it helps compare metal reactivity. Metals placed on hydrogen in this series can divert them from acids such as HCl and H2SO4 (because they are more reactive). Important Uses of Siri Aside from providing an understanding of the properties and reactivation of metals, the reactivation series has several other important applications. For example, the results of the between metals and water, metals and acids, and a single reaction between metals can be foreseen with the help of a series of activities. Reaction Between Metals and Calcium Water and metals that are more reactive than calcium in the reactivation series can respond with cool water to form a commensurate hydroxide while freeing hydrogen gas. For example, the response between potassium and the result of potassium hydroxide water and H2 gas, as explained by the chemical equations provided below. 2K + 2H2O → 2KOH + H2 Therefore, a series of metal dependencies can be used to predict the response between metal and water. The reaction between metal and plumbum acid and the position of the metal above precedes on the series of activities to form salts when reacting with hydrochloonic acid or sulfuric acid. This reaction also involves the release of hydrogen gas. The reaction between zink and sulfuric acid is an example of such a response. Here, zink sulfate and H2 gas are formed as products. The chemical equation is: Zn + H2SO4 → ZnSO4 + H2 Therefore, the response between the metal and some acids can be foreseen with the help of a series of reactivations. The reaction of a single piece of low-ranking metal ion metal is easily reduced by high-ranking metals on a series of reactivations. Therefore, low-ranking metals are easily moved by high-ranking metals in a single reaction between them. A great example of such a reaction is the copper ingest from copper sulfate by zinc. The chemical equation for this reaction is given by: Zn(s) + CuSO4 (aq) → ZnSO4 (aq) + Cu(s) This concept has several practical applications in metal extraction. For example, titanium is extracted from titanium tetrachloride through a single anjakan reaction with magnesium. Therefore, a series of metal shocks can also be used to predict the results of a single reaction. In the order of reactivity of the series the freeze indicates metal. The activeness of the metal can be attempted by examining its reaction in both reactions to competition and efforts. Elements in the periodic schedule known as transitional metals are less reactive, and metals such as gold and platinum support the bottom of the list, exhibiting less chemical response methods with any ordinary reagent. Jujukan metal reactivity, also known as the operating set, refers to the organization of metals in their ascending order of reactiveity. ... Metals tend to lose electrons and form kasi quickly. Some lead to the formation of metal oxides with ambient oxygen. The most reactive set of principles is alkaline metals (located far apart from mid metals and precious gases). Cesium is the second of the lower part of this group, has 6 electron grips, and it is commensurate with the characteristics of reactive atoms, making it the most reactive element. Non-metallic properties boiling point is rather low, and another metal is gas. Likewise, non-metallic metals are weak heat conductors, and non-picky metals are dull and brittle. Many are not highly reactive metals, while others do not respond at all. It depends on the number of electrons in the amount of outside energy. To find out more about this concept and other related concepts, such as electrochemical series, sign up with BYJU and download mobile applications on your smartphone. Text Book Institute CBSE NCERT Class 10 Subject Science (Chemistry) Name Ph sample Category 10 Science Lab Manual Reactivation Siri Metals Class 10 Introduction To Reactivation of elements > Atomic elements in the periodic schedule act in response to each other to form a new form (except noble gas). Atoms respond because of their tendency to lose, gain or share electrons. > metal: Metal atoms will always lose electrons to form positive ions called kasi. If an atom easily loses electrons at chamber temperature it is said to be reactive metal rather than metal whose atoms do not easily lose electrons at chamber temperature. > Not Metal: The advantage atom is not the metal or electron stock to form a negative ion called anion, the atom is not a metal if it easily gets electrons it is said to be more reactive. > response: We can mistook the speed of the metal's response by allowing them to respond with water at room temperature and collect the obtained hydrogen gas. If the water temperature, the amount of metal and the amount of water stored continue, the level of reactivation can be calculated. Why the backlash? > To regulate metals in an increasing order or reduce their activeness one of the experiments commonly used is to check their levels of food. This is a technique that can be trusted to use, since aluminum is actually a reactive metal but caused by a protective aluminum oxide on it, the results may indicate that the level of aluminum activeness is slow. > therefore the reaction is used to regulate the metal in the reactivation series. Metal Anjakan Reaction > More reactive metals will divert (take place) less reactive metals in the settlement of metal salts. Non-Metallic Anjakan Reaction > This is the same as the reaction of the metal. In halogens case, more reactive instead of metal will divert the less reactive ones out of place. Metal anjakan > For example, in a given chemical reaction: > Copper (II) sulfate is blue. When iron nails are dyed in this solution, the response takes place and the blue color disappears and exchanges greenish and the iron metal is overtopped with brown pink copper metal. > More reactive metals will always divert less reactive metals. > If the metal is less reactive to the more reactive solution of metal salts, there will be no response. > For example, less iron than zink. Iron + Zink sulfate —> No response. > in reaction, metals compete for anion instead of metal. Here, in the above reaction it is S042-ion. > Metal arrangements in the reactivation series can be attempted using this type of response. Siri > Metal arrangements in the order of their activeness exais called the activeness series or metal activity series. Mnemonic Element Symbol Of Reactivation Putting Potassium K As you can see this metal (excluding carbon, since carbon is used for metal reduction from metal oxide ore, all metals under carbon are subtracted by it) is on top of hydrogen in a series of reactivation so that they respond with acids and displace hydrogen gas. Metals + Acids -> Metal Salts + Hydrogen Some Nadium Na Light Lithium Li Color Calcium Ca Metal Magnesium Mg Around Aluminum Al Carbon Carbon C Zone Zinc Zn In Iron Fe The Tin Sn Light Lead Pb Here Hydrogen H H+ ions are responsible for its acidic properties. Come Copper Cu These elements are under hydrogen so that they do not respond with acid. (Acid contains H+ ions) Exception: Copper responds with concentrated nitric acid, nitrate ions oxidize copper. Metal Mercury Hg Silver Ag Gold Au Platinum Platinum Pt Note: When metals respond with nitric acid they do not release hydrogen gas (except magnesium and manganese) because nitric acid acts as a strong oxidizing agents. Siri Reactivation Class 10 Experiment 4 Goal (i) To observe the actions of Zn, Fe, Cu and Al metals on the following salt solutions: (a) ZnSO4(aq.) (b) FeSO4(aq.) (c) CuSO4 (aq.) (d) Al2(SO4)3(aq.) (ii) Arrange Zn, Fe, Cu and Al metals in a reduced renewal order based on the above decisions. Theory > Originally: Metal arrangements in order reduce their activeness called the reactivation series. The most reactive metals divert less reactive metals from their saline finish. > Colored metal salts > metal reaction with Salt Settlement (i) CuSO4 salt solution. As Al, Zn and Fe were able to divert Cu from his salt solution, all three of which were more reactive than Cu. (ii) FeSO4 salt settlement. Since Zn and Al were able to divert Fe from its salt solution thus, Zn and A1 were more reactive than Fe but Cu was less reactive than Fe. (iii) ZnSO4 salt solution. Since only Al metal diverts Zn from its saline solution thus, Al is the most reactive metal among Al, Zn, Fe and Cu. Al > Zn > Fe > Cu. Materials Require Four clean exam tubes, markers, a piece of sand paper and the establishment of an exam tube. Necessary chemicals: Aluminum sulfate settlement, copper sulfate settlement, zink sulfate settlement, iron sulfate settlement, iron metal paths, zinc; copper and Procedure > Reaction with CuSO4 solution: 1. Take four clean test tubes. 2. With markers labeling them as A, B, C and D. 3. Take a copper sulfate solution in each test tube. 4. Swim a piece of aluminum, zinc, small and clean, clean, and copper metal in test tubes A, B, C and D respectively. 5. Record your observations. Conclusion: Metals Al, Zn and Fe are more reactive than Cu. Hence, Cu is the least reactive metal among the four given metals. > with FeSO4 solution: 1. Take four clean test tubes, labeled as A, B, C, and D. 2. Take the FeSO4 (aq) solution in each test tube. 3. Small dips, clean pieces of aluminum, zinc, iron and copper metal in their respective test tubes A, B, C and D. 4. Record your observations. Conclusion: Metals Al and Zn are more reactive than Fe and Cu. > Reaction with ZnSO4 solution: 1. Take a test tube A, B, C, and D, clean it up. 2. Add a ZnSO4 solution in each test tube. 3. Small dips, clean pieces of Al, Zn, Fe and Cu metal in test tubes A, B, C and D respectively. 4. Record your observations. Conclusion: Metal al is more reactive than Zn. > metal with Al2 (SO4)3: 1 solution. Take the test tubes A, B, C, and D and clean them. 2. Add Al2(SO4)3 solution in each test tube. 3. Small dips, clean pieces of Al, Zn, Fe and Cu metal in test tubes A, B, C and D respectively. 4. Record your observations. Conclusion: Al metal is not transferred by any given metals namely Al, Zn, Fe and Cu. Hence, Al Metal is the most reactive metal among the given metals. 1. Al is not moved by any of the four metals from his salt solution, therefore, Al is at the top of a series of reactivity. 2. Al can reckon Zn from the salt solution but no other metal can reckon it. Hence, the activation order is Al > Zn. 3. Metal Al and Zn can distancing Fe metal from salt solutions but Cu cannot. Hence, the metal order in order of reduced reactivity is Al > Zn > Fe > Cu. I. CuSO4 in each test tube. II. FeSO4 Solutions in each test tube. III. ZnSO4 Solutions in each test tube. IV. Al2(SO4)3 solution in each test tube. Precautionary Steps 1. Clean the metal by rubbing it with a piece of sandpaper before dipping it in a salt solution. 2. Wash the test tube after each set of observations of certain metal interactions with four salt solutions. 3. Use a little copper sulfate saturation solution, aluminum sulfate, iron sulfate and zinc sulfate. 4. Use very small pieces of metal each time. 5. Do not touch any chemicals. Viva Voce Question 1: What is the color of the copper sulfate solution? Answer: The color of the copper sulfate solution is blue. Question 2: What is the color of the iron sulfate solution? Answer: The color of the iron sulfate solution is green. Question 3: When we added Fe in the FeSO4 solution, no changes were seen. Why? Answer: It is because Fe did not respond with the FeSO4 solution. Question 4: one metal which when added to the settlement of cuso4 blue converts it to green. Answer: The iron metal responds with a blue copper sulfate finish and turns it green. Question 5: What is the finishing color of Al2(SO4)3? Al2(SO4)3? The al2(SO4)3 solution is colorless. Question 6: Complete the chemical equation given Fe(s) + CuSO4(aq) —answer >: Fe(s) + CuSO4(aq) —> FeSO4(aq) + Practical Questions Cu Question 1: What is a series of metal activity? Answer: The order of metal in order reduces their reactivity called a series of metal activity. Question 2: What is the dissensive reaction? Answer: When a more reactive metal diverts less reactive metal from its salt solution, it is called a displacing reaction. Question 3: What happens when a piece of zinc metal is introduced into a newly prepared solution ferus sulfate? Answer: The color of the ferous sulfate solution is green and when the zinc is added to it the color fades and becomes colorless. Question 4: What happens when metal Fe (Iron) is added to the solution of blue copper sulfate? Answer: Iron metal reacts with a solution of copper sulfate to divert copper and form a green iron sulfate. Fe(s) + CuSO4(aq) —> FeSO4(aq) + Cu(s) Question 5: What happens when A1 metal is added to a copper sulfate solution? Answer: Al metal reacts with a solution of copper sulfate to form aluminum sulfate and copper metal evacuated. Question 6: Name the least reactive metal between the following: Fe, Zn, Al. Answer: Fe is the least reactive metal among Fe, Zn, and AT Question 7: What will happen if Fe metal is added to the ZnSO4 solution? Answer: No reaction will occur, because metal fe is less reactive than Zn. Question 8: The metal arrangements given in the series of reactivity: Mg, Zn, Cu, Fe, Al Answer: The metal reactivity series is Mg > Al > Zn > Fe > Grand Question 9: Name some metal that will respond with an aluminum sulfate solution, that is, Al2(SO4)3. Answer: Metals such as calcium, sodium and magnesium will react with a sulfate aluminum solution. Question 10: When you keep an aluminum strip in a FeSO4 solution, what changes are seen? Answer: Aluminium is more reactive than Fe therefore it diverts Fe from FeSO4 and the green color of FeSO4 disappears. Question 11: Why is Fe metal more reactive than copper metal? Answer: Fe easily loses electrons versus copper. Therefore, it is more reactive than copper metal. Question 12: How can you test that the sample given contains water or not? Answer: On the heating of the sulfate sulfate crystals of copper blue slowly change to light blue and then colored. Water drops are collected on the inside of the test tube. Question 14: It is considered that each molecule of copper crystals at room temperature contains five water molecules as crystalline water. Do you see any difference in them? (Hint: See the reaction of copper sulfate dehydration) Answer: one copper sulfate molecule, five water molecules present to form crystals. The blue color of the crystal is caused by these five water molecules. But, in anhydrous copper anhydrous water molecules are not present. NCERT Makmal Manual Question 1: Why does the color of copper sulfate solution change, when iron nails are dyed in it? Answer: Iron nails respond with copper sulfate and form iron sulfate. Due to the reaction of the blue color of copper sulfate changes to the green color of ferrous sulfate. Question 2: How will you summarize the procedure to show that Mg > Fe > Cu is in a series of reactivations? What are the basic principles involved in this Experiment? Answer: I will respond to each metal (Mg, Fe and Cu) with a saline solution of Mg, Fe and Cu. The metal that will divert the two metal ions from the salt solution will be the most reactive and the one that will not show any discoloration in the solution is the least reactive. Principle: The most reactive metals will divert less reactive metals. Question 3: Why does the following reaction apply? 2I- (aq) + Cl2 (aq) —> 2I– (aq) + Answer I2 (resolved) : This is a reaction seen among halogens. Chlorine is more reactive than ioin and therefore it diverts iodide ions. Question 4: In the following reactions, A and B are metals. BX is a metal salt B. A + BX —> AX + B Which one of the two metals is more reactive? Give me a reason. Answer: Metal A is more reactive than B, as metal A diverts metal B from its salt solution. Question 5: Name any of the two metals that are more reactive than iron. Answer: Zink and magnesium. Question 6: Why does the finish color of copper sulfate (II), change, when the zink metal decreases in it? Answer: Zink is more reactive than copper and diverts copper from its settlement. Thus, the blue color of copper sulfate settlement becomes colorless. Question 7: What is your observation when copper is added in the solution of ferrous sulfate (II) ? Answer: Copper is less reactive than iron and should not divert iron from iron sulfate solutions and no discoloration. Question 8: Why can we keep iron (II) safely in copper vessels whereas the same cannot be kept safely in zink ships? Answer: Copper is less reactive than iron and should not respond with ferrous sulfate. While zink is more reactive than iron and therefore it can respond with ferrous sulfate. Various Choice Questions (MCQ) Questions based on Procedure Proficiency and Manipulative 1. When the iron line is placed in the settlement of copper sulfate, the time required for the settlement color to change from blue to green will be less if: (a) it is cooled (b) it is shaken (c) it is kept undisturbed (d) it is heated. 2. The finish color of the sulfate zink is (a) white (b) green (c) yellow (d) colorless. 3. Will the following reactions not apply? (a) Al + ZnSO4 (b) Zn + (c) Cu + ZnSO4 (d) Mg + ZnSO4. 4. What are the following reactions that will apply? (a) Fe + FeSO4 (b) Cu + FeSO4 (c) Au + FeSO4 (d) Mg + ZnSO4. 5. To show the show zink is more reactive than copper, the correct procedure is to (a) provide a solution of copper sulfate and stick a zink line in it. (b) provide a settlement of copper sulfate and stick copper lines in it. (c) heat zink and copper lines. (d) add liquid nitric acid on both pathways. 6. When adding iron lines in the settlement of copper sulfate, the response takes a long time, to make it happen quickly we can: (a) add more iron lines (b) add more settlement of copper sulfate (c) heat the traces of solids (d) iron heat and add in copper sulfate. 7. You are given two exam tubes A and B, with the completion of ZnS04 and A12(S04)3, both colorless. To identify a solution, (a) add Zn in exam tubes A and B (b) plus A1 in exam tubes A and B (c) first add Zn and then A1 in exam tubes A and B (d) all of the above are correct. 8. Two metals that can keep iron away from its solution are: (a) copper and zink (b) zinc and tin (c) silver and aluminum (d) zinc and aluminum. 9. P+QR → PQ+R. In the above reaction, (a) Q displaces P(b) R is moved by P(c)P displaces Q(d)R displaces Q. 10. Two A and B defects contain iron sulfate solution (II). In the bearer A is placed a small piece of copper and inside the B-strangler is placed a small piece of zink. It is found that the form of a gray deposit on the zink but not in copper. From this observation it can be concluded. (a) Zinc is the most active metal followed by iron and copper. (b) Zinc is the most active metal followed by copper and then iron. (c) Iron is the most active metal followed by zinc and then copper. (d) Iron is the most active metal followed by copper and then zinc. Questions based on ObservationAl Proficiency 11. A student takes the Cu, Fe, Zn and Al lines separately in four exam tubes labeled I, II, III and IV. He added 10 mL of newly prepared sulfate ferus completion for each test tube as shown below: Black remains will be obtained in the test tubes: (a) I and II (b) I and III (c) II and III (d) III and IV. 12. Aluminum powder added to the settlement of copper sulfate. The color of the settlement changes from: (a) not blue (b) blue to (c) light green to blue (d) reddish brown to light green. 13. The color of the coriander leaves obtained on the zink stem on the dye in the settlement of aqueous copper sulfate will be (a) blue (b) brown (c) white (d) green. 14. Which is not noticed when aluminum is added to the copper sulfate solution? (a) The solution is blue at first. (b) The final settlement becomes colorless. (c) The final settlement becomes lightly green. (d) Brown jisim on the surface of aluminum. 15. Copper turn added to the solution of salt without color. After 10 minutes it was noticed that the colorless solution switched to a colored solution and a flashing metal immersion appeared. The finishing color and metal will each be (a) yellow and Zn(b) green and Fe(c) red red Sn(d) blue and Ag. 16. The filing of iron has been added to the settlement of copper aqueous sulfate. After some time, at the observation, it was found that the color of the solution had changed from (a) blue to pale green (b) blue to dark green (c) blue to (d) blue to reddish-brown. 17. Zink granule is added to zink sulfate, copper sulfate, aluminum sulfate and iron sulfate finish as shown below. You will notice metal lings on the zink in the test tubes (a) I and III (b) II and IV (c) I and II (d) III and IV. 18. A student adds a piece of zink metal in four different exam tubes containing different solutions. Where's the exam tube he didn't notice no change? (a) Test the tube with the completion of the CuSO4 (b) Tube exam with the completion of the AgNO3 (c) Tube exam with the completion of the ZnSO4 (d) Tube exam with the completion of FeSO4. 19. A student conducts the following four experiments: He will find the formation of metal steeping in experiments (a) II, III (b) I, III (c) I, III (d) II, III, IV. 20. A piece of granulated zink has been dropped into the settlement of copper sulfate. After a while, the color of the solution changes from (a) light green to blue (b) blue to (c) green to (d) blue to yellow. 21. Copper lines are placed inside the hoe containing the completion of sulphate zink. When observing the path the next day, it is realized that (a) the copper line remains because it (b) the copper line becomes thinner (c) the copper line becomes thicker (d) the color of the path changes. 22. If some copper sulfate crystals are dissolved in water, the finish color obtained is: (a) green (b) red (c) blue (d) brown. 23. The filing of iron has been added to the settlement of copper sulfate. After 10 minutes, it is noted that the blue color changes the solution and the layer will be deposited on the iron filing. The finish color and coriander will each be (a) yellow and green (b) brown and blue (c) red and greenish-blue (d) green and reddish-brown. 24. Iron sudu is placed inside the glass marks that hold the sulfate zink finish. What is the correct observation? (a) Zink sulfate will be green. (b) The glass marks will get cracked. (c) Iron sudu will get a hole in it. (d) Iron sudu will not show any changes. 25. If the stored aluminum line is immersed in the completion of the newly supplied ferus sulfate taken in the test tube, the noticed change is (a) the green solution slowly swapping blue. (b) the lower end of the test tube becomes slightly warm. (c) colorless gases with a burning sulfur smell are noticed. (d) the light green finish turns blue. 26. Completion of FeSO4, ZnSO4, CuSO4 and Al2(SO4)3 settlements has been separately in four test tubes and several iron nails are placed in each solution. After a few minutes, it will be noted that the color (a) of the four solutions changes. (b) ZnSO4, ZnSO4 solution, and Al2(SO4)3 changed and FeSO4 was unchanged. (c) the solutions of ZnSO4 and Al2(SO4)3 are changed only. (d) Copper sulfate solution only changed. 27. When you put the iron nails in the solution of copper sulfate, the reddish brown coat that forms on the nails is (a) soft and numb. (b) hard and not fragmented. (c) smooth and shine (d) rough and granular. Questions based on Reporting and Interpretation Skills 28. The 10 mL of newly prepared iron sulfate has been taken in each of the four test tubes. Copper, iron, zinc and aluminium strips are introduced, each of the metals in different test tubes. One black waste is obtained in two of them. The right pair of metals forming precipitates are (a) Copper and zinc (b) aluminum and copper (c) iron and aluminum (d) zinc and aluminum 29. Four students were assigned separately to the experiments of iron nail interaction with copper sulfate solutions. Each group records observations as given below in the table. Which group of students recorded proper observations? 30. A student takes four test tubes containing different color solutions marked I, II, III, and IV as shown below. Test tubes containing copper sulfate solutions and ferous sulfate solutions can be tubes: 31. Four A, B, C and D students specify the initial color of the settlement in testing tubes I. II, III and IV. After silent the zinc stick in each solution and leaving it undeterred for two hours, they print the color of each solution again. They recorded their observations in the form of a table given below: Students Colorless Solution I II III IV Early Colored Green Blue Final ColorLess B Red Light Lights Red Light Green ColorLess Light ColorLess C Light ColorLess Red Light Blue Light Blue Blue Light Blue Light Blue Light Blue Light (a) A (b) B (c) C (d) D. 32. Aluminium sulfate and copper sulfate solutions were taken in two test tubes I and II. Several pieces of iron filing are then added to both solutions. All four students A, B, C and D record their observations in the form of tables as given below. Students (I) Al2(SO4)3 Solutions (II) Copper Sulfate Solution Change Solutions Without Color to the blue blue green color solution is maintained. B The colorless solution remains unchanged for change of Blue solution to green. C Changes the colorless solution to blue light of blue color changes to the green light. D Colorless solutions remain unchanged Blue Color faded solutions. The correct set of observations was recorded by the student (a) A (b) (c) C (d) D. Scoring Keys With Explosion 1. (d) Heat increases the response rate. 2. 2. The solution of zink salt is colorless. 3. (c) Cu is less active than Zn. 4. (d) Mg is more reactive than Zn. 5. (a) This will show a backlash. 6. (c) Increased temperature speed resulting in retaliation. 7. (d) All exams will assist in the exact completion. 8. (d) Both are more reactive than Fe. 9. (b) R is removed from the compound and exists freely. 10. (a) Copper does not respond, but zink responds with the solution of ferus sulfate. Therefore, zink is more reactive followed by iron and copper. 11. (d) Zn and Al are more reactive than Fe and may divert them. 12. (b) Copper salt is blue and aluminum salt is colorless. 13. (b) Cu is moved by Zn who is brown. 14. (c) Aluminum salt is colorless and not green in color. 15. (d) Copper is more reactive than Ag. 16. (a) Fe keep copper ions away. 17. (b) Zink is more reactive than Cu and Fe. 18. (c) Equilibrium views. 19. (c) According to the series of activeness. 20. (b) The solution of Zink salt is colorless. 21. (a) Copper is less reactive than zinc. 22. (c) Salt cu is blue. 23. (d) Ferrous salt is green. 24. (d) Iron is less active than Zn. 25. (b) It shows a terrorist reaction and heat develops. 26. (d) Cu is the least reactive metal and can be moved by all given metals. (a) Newly deposited copper is soft and dull. 28. (d) More reactive zink and alufninium will replace iron rather than ferrous sulfate. 29. (d) Fe will put cu ions to give a settlement of green color and brown-ignited Cu metal. 30. (d) Copper sulfate is blue and ferus is green in color. 31. (a) In the first two test tubes, the metal will not retaliate but in the other two it will retaliate. (b) Fe is less reactive than Al but more reactive than Cu. Therefore, the backlash will prevail. We hope this CBSE Class 10 Science Lab Reactivation Series helps you in your preparation for the CBSE Class 10 Institute Examination. For any questions regarding CBSE's Class 10 Practical Science Siri Materials, please ask questions in the comment room. Be careful to pack the current current information about NCERT Books, NCERT Completion, CBSE Sample Paper, Aggarwal RS Completion and all the sources we share on a regular CBSETuts.com for K12 Students. More CBSE Source Class 10 Practical Proficiency Manual Makmal: Proficiency: