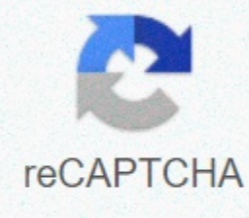




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## State the periodic law and explain how it was arrived at

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By 1829, chemist Johann Wolfgang Dobreiner noted that some elements with similar characteristics occurred in a group of three such; Chlorine, bromine and iodine; calcium, strontium, barium; sulfur, selenium, tellorium; iron, cobalt and manganese. However, at the time of this discovery very few elements were discovered and there was a confusion between molecular weight and atomic weights; Therefore, chemists did not really understand the importance of the Dobner Triad. In 1859, two physicists, Robert Wilhelm Ponson and Gustav Robert Kerchove, discovered the spectroscopy that allowed the discovery of many new elements. This gave scientists the tools to detect relationships between the elements. Thus, in 1864, chemist John R. Newland arranged the elements in the increase of atomic weights. Explaining that a certain set of characteristics is repeated every eight places, he is named the Act of Octaves. In 1869, Dmitry Mendeleev and Lothar Meyer came individually with their own periodic law when the elements are arranged in order to increase the atomic mass, certain sets of characteristics repeated periodically. Mayer based its laws on the size of the offspring (the atomic mass of an element divided by the intensity of its steel form), this property is called the size of the mueller.  $\rho = \frac{M}{V}$  Leave empty spaces for undiscovered items (atomic blocks, element: 44, scandium; 68, gallium; 72, Germanium;&100, technetium) so that certain elements can be grouped together. However, Mendelife did not expect noble gases, so they were left no stains. Figure 1: The original periodic table of Mendelife in the Mendeliev table, items with similar characteristics are located in portrait Groups are called groups. The high size of the mall from top to bottom group3 example of alkaline minerals (Mendeleev I Group) has high amounts of sire and they also have low melting points that go down in order: me (174 cc) &gt; Na (4&gt; K (63.7 oC) &gt; Rb (38.9 x X) &gt; C (28.5 o C) assuming errors in atomic blocks, mendelife put certain elements not in the order of increasing the atomic mass so that they can fall into the appropriate groups (similar elements with similar characteristics) of his table Periodic. An example of this was with Argon (atomic mass 39.9), which was placed in front of potassium (atomic mass 39.1). The elements were placed in groups that expressed similar chemical behavior. In 1913 Henry J.J. Mosley researched the X-ray spectra of the elements and suggested that the energies of electron orbits depended on the nuclear charge and the nuclear charges of atoms in the target, also known as anode, dictate the emitted X-ray frequencies. Mosley was able to link X-ray frequencies to figures equal to nuclear charges, thus showing the placement of the elements in the Mandlev periodic table. The equation he used:  $\nu = A(Z-b)^2$  with  $\nu$ : X-Ray frequency  $\nu$ (Z): Atomic number  $\nu$ (A) and  $\nu$ (b): Constants with the contribution of the periodic law mousley can be reaffirmed: similar characteristics are repeated periodically when the elements are arranged according to the increase in the number of atoms. Atomic numbers, not weights, determine the chemical properties factor. As mentioned before, Argon weighs more than potassium (39.9 vs. 39.1, respectively), but argon is in front of potassium. Thus, we can see that the elements are arranged on the basis of their atomic number. It has been found that periodic law helps to identify many patterns of different characteristics of the elements; The table below shows that the elements increase from left to right accordingly to their atomic number. Vertical columns have similar properties within their group for example lithium-like sodium, beryllium resembles magnesium, and so on. Group 1 2 13 14 15 16 17 18 Element Lee to be B C N O F Ne Atomic No. 3 4 5 6 6 7 8 9 10 10.81 12.01 14.01 15.99 18.99 20.18 Element na Mg Al Si P Cl Ar Atomic Number 11 13 14 15 16 17 18 Atomic Mass 22.99 24.31 26.98 20.09 30.97 32.07 35.45 39.95 for elements in group 1 (periodic table) with similar chemical properties called metals Alkaline. The elements in group 2 have similar chemical properties, called alkaline ground metals. The periodic table of the short shape is a table where the elements are arranged in 7 rows, periods, with increasing atomic numbers from left to right. There are 18 vertical columns known as groups. This table is based on the periodic table of Mendeleev and the periodic law. In A long form, each period is related to the construction of an electronic shell; The elements of the group are called 18 noble gases, and the group 17 is called halogens. F-block elements, called internal transport minerals, which are located at the bottom of the periodic table (periods 8 and 9); References Torushi, Ralph H., William S. Harwood, F. G. Herring, and Jeffrey D. Madura. General Chemistry: Modern Principles and Applications. 9th ed. River Saddle High: Pearson Education, Inc., 2007. Sisler, Harry H. Electronic Structure, Property, and Periodic Law. New York; New York, Reinhold Publishing, 1963. Petrucci, Ralph H., Carrie Bissonnette, F.J. Herring, and Jeffrey D. Madura. General Chemistry: Modern Principles and Applications. Custom version for CHEM 2. Pearson Learning Solutions, 2010. Mendelife's periodic schedule in 1869. At Wikimedia Commons. Viewed on December 4, 2010, from commons.wikimedia.org/wiki/File:..odic\_table.png. Table of periodic items. At Wikimedia Commons. Retrieved December 4, 2010, from commons.wikimedia.org/wiki/Pe...le\_of\_Elements. 1) Periodic law states that similar properties are periodically repeated when items are arranged according to the number of atomic properties repeated periodically when the elements are arranged according to atomic weight weight similar properties everywhere on the elements of the periodic table in the same period have the same characteristics 2) Which element is the most similar sodium potassium aluminum calcium oxygen 3) according to periodic law, will argon be in front of potassium or after? Explain why 4) What is the most calcium-like element? Carbon oxygen strontium iodine 5) who were the two chemists that came with the periodic law? John Dalton, Michael Faraday, Dimitri Mendeleev, Lothar Meyer, Michael Faraday, Lothar Meyer, John Dalton and Dmitri Mendeleev answers a argonaut suppin because the periodic law stipulates that the periodic table increases from left to right based on atomic number and not atomic weights C B educational goals state law league. Description of the organization of the periodic table. We all enjoyed playing an idea game. The goal of the game is to get information about a murder - who did it, where they did it, and what was used as a murder weapon. As the game progresses, clues are obtained by each player, and they must then group these clues into guessing as to the criminal. An individual piece of information takes on a broader significance when placed along with other parts of the puzzle. When Mendelife his periodic table together, no one knows about the existence of a nucleus. It was not until 1911 that Rutherford conducted his gold foil experiment that showed the presence of the nucleus in the atom. Just two years later, in 1913, The English physicist Henry Mosley (1887-1915) studied x-ray spectra for a number of chemical elements. He released X-rays through element crystals and studied the wavelengths of the radiation he discovered. Mosley found that there was a relationship between wavelength and atomic number, the results of which defined the atomic number as the number of protons in the nucleus of each atom. He then realized that the elements of the periodic table should be arranged by arranging an increase in the atomic number rather than the increase of the atomic mass. When ordered by an atomic number, the contradictions within mendelifev's table disappeared. Tellorium has an atomic count of 52, while iodine has an atomic number of 53. Even though tellorium actually has the largest atomic mass of iodine, it is properly placed before iodine in the periodic table. Mendeleev and Mosley are credited with

the greater responsibility for modern periodic law: when the elements are arranged in order to increase the atomic number, there is a periodic repetition of their chemical and physical properties. The result is the periodic table as we know it today. Each new horizontal row of the periodic table corresponds to the beginning of a new period because the new main energy level is filled with electrons. Elements with similar chemical properties appear at regular intervals, within vertical columns called groups. Summary elements of the periodic table are arranged in the order of an atomic number increase. The periodic law stipulates that when elements are arranged in order to increase the atomic number, there is a periodic repetition of their chemical and physical properties. Use the link below to answer the following questions: where Mosley went to university? Who did he research with after he graduated from college? What is Mosley's law? A review. Who discovered the relationship between x-ray wavelength and atomic number? What did Mosley come up with from his research? What is periodic law? What do vertical columns (groups) represent in the periodic table? Group: Items with similar chemical properties appear at regular intervals, within vertical columns. Period: The period is a horizontal row of the periodic table. Periodic law: When the elements are arranged in order to increase the number of atoms, there is a periodic repetition of their chemical and physical properties. Properties.

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