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## Chemistry chapter 5 section 2 review answers

The elements in the s and p blocks of the periodic table, called elements in group 18, have nonmetallic elements in the group of 17 that respond to metals, mainly in salt form, namely the outer electron configuration of a neutral alkaline metal atom consisting of two electrons in orbit in orbit. Naturally, alkaline metals are found only in compounds, because they are very reactive elements, where the hydrogen group is an element found in groups of 3-12 of the periodic table, classified as period figures, indicating that the maximum energy level is complemented by the characteristics of silver alkaline metals, the appearance of softness. The very reactive characteristics of alkaline metals, the soil is harder, denser and stronger than alkaline metals. Block d elements are metal with general metal properties, characteristics of the composition, changes, good conductors of boron, electric fluorescent, germanium silicon. Arsenic, antimony, and halogen tellurium reacted vigorously with \_\_\_\_\_ to create a compound called \_\_\_\_\_ characteristic of shiny lanthanides, similar metals in reaction to groups of 2 elements in the s and p groups of the periodic table, called elements in group 18. In orbits in nature, alkaline metals are found only in compounds, because they are very reactive elements that the hydrogen group is an element found in groups 3-12 of the periodic table, classified as period figures indicating the maximum energy level filled with the characteristics of silver alkaline metals. Softness is not found in nature as a free element, the very reactive characteristics of alkaline metals, more difficult, denser and stronger than alkaline metals. The higher the melting point, the less reactive than the alkaline metal, the excessive reaction to be found in nature. The d block element is a metal that resembles the general metal properties of the composition, changes the good conductors of boron, electric glow, silicon, germanium, arsenic, antimony and halogen tellurium, reacting vigorously with \_\_\_\_\_ to create a compound called \_\_\_\_\_ characteristics of metal, similar lanthanides shadows in reaction to this group of 2 metal Ch 5 chemistry (Pearson) Ch 5 electrons in atoms. Places in atoms where electrons tend to be most common. Quantum energy supply needed to move electrons from one energy level to another. Quantum Mechanical Model pearson chemistry with free interactive flash cards Choose from 500 different combinations. 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