



Le chatelier worksheet

In this worksheet, we will deal with using Le Chatelier's principle to predict directions of change for equilibria in response to stimulation. Q1: In equilibrium, a box contains five molecules each from gases A, B, and C. Shown in the chart are the equilibrium gas mixtures that are obtained when molecules of A or B are added to the box without changing the overall pressure. Write a balanced chemical equation for the reaction that occurs in the gas mixture is increasing and the system may reach equilibrium. Assuming the temperature remains at the original value and no molecules are added or removed, which of the following situations can be expected? B C D E molecules are added or removed from the original values throughout this process. Which of the following changes cannot cause the displayed state? AAdding B and C BAdding B and removing CAdding A and C DAdding A and C and removing Q2: Carbonate solid calcium breaks down thermally to create solid calcium breaks down thermally to create solid calcium oxide and co2 gas, With an enthalpy change of +192 kJ/mol. The equation for this response is as follows: CaCO()()+CO()A 32ssgA sample of thermally decaying carbon dioxide calcium at a constant temperature under the atmosphere of carbon dioxide in the constant pressure of one ATM, as is done in the chart. Once equilibrium has been achieved, calcium carbonate mixture and calcium oxide is heated to a higher temperature. How and why does the volume of the gas balloon change? The position of the equilibrium location does not change, so the storage space does not change. C The position of equilibrium moves to the left, increasing the volume. D The position of equilibrium moves to the right, reducing the volume. E The position of the equilibrium changes to 2 ATMs. How and why does the solid mass in the heated test tubes change? The position of the equilibrium changes to the left, increasing the solid mass. BThe equilibrium position of the equilibrium moves to the right, reducing solid mass. The position of the equilibrium does not change, so the solid block does not change. Once the equilibrium has reached its equilibrium, a small mass of calcium carbonate is added to the heated sequecon. How and why does the concentration. BThe position of equilibrium moves to the left, increasing the concentration. C The position of equilibrium moves to the right, increasing the concentration. D The position of equilibrium does not change, so the concentration is unchanged. The position of equilibrium changes to the right, reducing concentration is unchanged. The position of equilibrium changes to the right, reducing concentration is unchanged. equation: PCIPCI +CI532 The enthalpy change to this response is +87.9 kJ/mol. Does the B C D E diagram represent the rates of reactions back and forth when the pressure drops at a constant temperature? B C D E Q4: Colorless nitrogen gas tetroxide (no)24 invertedly decomposed into nitrogen dioxide heat gas (no)2 with enthalpy response of +5 7.5.2 2 kJ/mol: NO2NO242 A student stores a sample of a nitrogen tetoxid in a gas syringe at room temperature and presses one ATM and allows the syringe's plunger to move freely. The gas was left to be buried to equilibrium, as it showed. Why is the piston's position in equilibrium different from its position at the beginning of the response? The pressure of the gas in the syringe is higher. B The syringe contains gas molecules of a larger average size. The temperature of the gas in the syringe contains a greater number of gas molecules. The syringe contains a larger mass of gas molecules. The student heats the gas injector and allows the syringe's loch to move freely. Which diagram best illustrates the gas injecter in the new reaction conditions after equilibrium is achieved? B C D E The student compresses the gas by pushing the plunger and allows the reaction to reach equilibrium at room temperature, with the plunger fixed at its new location. Which diagram best illustrates the gas injecter in the new reaction conditions after equilibrium is achieved? B C D E Q5: Chlorometan reacts with sodium chloride, as on the other hand the equation. The grinding solubilities of sodium cloride and sodium conditions after equilibrium is achieved? B C D E Q5: Chlorometan reacts with sodium chloride, as on the other hand the equation. CAcetone DFormamide EDimethylformamide Q6: Under certain conditions, nitric oxide (no) and nitrogen dioxide (no) React inversely to produce liquid di nitrogen trioxide (no) React inversely to produce liquid di nitrogen trioxide (no) and illustrates the effects of temperature (T) and pressure on the equilibrium yield of this reaction? B C D Q7: In equilibrium, a box contains five molecules each from gas compounds A, B, C and D. Changing the temperature, T, or pressure, P, of the system changes the composition of the gas mixture in equilibrium, as shown in the chart. Write a balanced chemical equation for the reaction that occurs in the gas mixture. A2AB+C BA+2CD CA+CD DA+2CB+D E2A+2CB The temperature of the box is decreasing and the system may reach equilibrium. Assuming the pressure remains at the original value, which of the following situations can be expected? B C D E increased pressure and the system may reach equilibrium. Assuming the temperature remains at the original value, which of the following situations can be expected? B C D E Q8: יונים, כפי שמופנת על-ידי המשוואה הבאה [Cr((OH)(OH)()+3OH()]23] מגיב עם יונים הידרוקסיד עודף במים כדי לייצר מים (Cr((OH)(OH)()+3OH())[Cr((OH)()+3OH()]23] מגיב עם יונים הידרוקסיד עודף במים כדי לייצר מים (Cr((OH)(OH)()+3OH())[Cr((OH)()+3OH()]23] מגיב עם יונים הידרוקסיד עודף במים כדי לייצר מים (Cr((OH)(OH)()+3OH())[Cr((OH)()+3OH()]23] מגיב עם יונים הידרוקסיד עודף במים כדי לייצר מים (Cr((OH)()+3OH())[Cr((OH)()]23] מגיב עם יונים הידרוקסיד עודף במים כדי לייצר מים (Cr((OH)())[Cr((OH)()]23] מגיב עם יונים הידרוקסיד עודף במים כדי לייצר מים (Cr((OH)())[Cr((OH)()]23] מגיב עם יונים הידרוקסיד עודף במים כדי לייצר מים (Cr((OH)())[Cr((OH)()]23] מגיב עם יונים הידרוקסיד עודף במים כדי לייצר מים (Cr((OH)())[Cr((OH)()]23] יונים, כפי שמוצג – aqueous [Cr(OH)] מיוצר על ידי הוספת X הם ירוק כחול בצבע ולהכהות עם ריכוז הולך וגדל. כל היונים המים האחרים בתגובה זו מייצרים פתרונות חסרי צבע. השעיה – 3[(Cr(HO)(OH)] מיוצר על ידי הוספת X הם ירוק כחול בצבע ולהכהות עם ריכוז הולך וגדל. כל היונים המים האחרים בתגובה זו מייצרים פתרונות חסרי צבע. השעיה – 3[(Cr(HO)(OH)] מיוצר על ידי הוספת A הם ירוק כחול בצבע ולהכהות עם ריכוז הולך וגדל. כל היונים המים האחרים בתגובה זו מייצרים פתרונות חסרי צבע. השעיה – 3[(Cr(HO)(OH)] מיוצר על ידי הוספת A הם ירוק כחול בצבע ולהכהות עם ריכוז הולך וגדל. כל היונים המים האחרים בתגובה זו מייצרים פתרונות חסרי צבע. השעיה – 3[(Cr(HO)(OH)] מיוצר על ידי הוספת A הם ירוק כחול בצבע ולהכהות עם ריכוז הולך וגדל. כל היונים המים האחרים בתגובה זו מייצרים פתרונות חסרי צבע. השעיה – 3[(Cr(HO)(OH)] מיוצר על ידי הוספת A הם ירוק כחול בצבע ולהכהות עם ריכוז הולך וגדל. כל היונים המים האחרים בתגובה זו מייצרים פתרונות חסרי צבע. השעיה – 3[(Cr(HO)(OH)] מיוצר על ידי הוספת A הם ירוק כחול בצבע ולהכהות עם ריכוז הולך וגדל. כל היונים המים האחרים בתגובה זו מייצרים פתרונות חסרי צבע. השעיה – 3[(Cr(HO)(OH)] מיוצר על ידי הוספת A הם ירוק כחול בצבע ולהכהות עם ריכוז הולך וגדל. כל היונים המים האחרים בתגובה זו מייצרים פתרונות חסרי צבע. השעיה – 3[(Cr 63 - Ions and BRemoving Acid [Cr(OH)]63 - Ions and Adding Base E 63- Ions and Adding Base E 63- Ions and Adding CAdding Acid [Cr(OH)]63 - Ions and Adding CAdding Acid [Cr(OH)]63 - Ions and Adding Base E 63- Ions and Adding CAdding Acid [Cr(OH)]63 - Ions and Adding Base E 63- Ions and Adding CAdding Acid [Cr(OH)]63 - Ions and Adding CAdding Base E 63- Ions Adding CAdding Acid [Cr(OH)]63 - Ions and Adding Base E 63- Ions Adding CAdding Acid [Cr(OH)]63 - Ions and Adding Base E 63- Ions Adding Base E 63- Ions Adding CAdding Acid [Cr(OH)]63 - Ions Adding Base E 63- Ions Adding Acid [Cr(OH)]63 - Ions Adding Base E 63- Ions Adding Base E Homework and Homework Handouts Need Help Downloading: Doc File: You Need the Microsoft Word Displays, or a program that can import Word files to view that file. To learn more about the free Microsoft Word Viewer, visit the Microsoft Word Web site. Mr Abraham's Site Penfield High School 25 Penfield Drive High School, NY 14526 In this worksheet, we will deal with using Le Principle to predict directions of change for equilibrium, a box contains five molecules each from gases A, B, and C. Shown in the chart are the equilibrium gas mixtures that are obtained when molecules of A or B are added to the box without changing the overall pressure. Write a balanced chemical equation for the reaction that occurs in the gas mixture. AA+B2C BA+2C2B CA+BC DA+2CB EA+C2B The total pressure of the original gas mixture is increasing and the system may reach equilibrium. Assuming the temperature remains at the original value and no molecules are added or removed, which of the following situations can be expected? B C D E molecules are added or removed from the original gas mixture and the system may reach equilibrium. Temperature and pressure are fixed in their original values throughout this process. Which of the following changes cannot cause the displayed state? AAdding B and C BAdding B and C BAdding A and C DAdding A and C DAdding A and C and removing Q2: Carbonate solid calcium oxide and co2 gas, With an enthalpy change of +192 kJ/mol. The equation for this response is as follows: CaCO() ()+CO()A 32ssgA sample of thermally decaying carbon dioxide in the air. Product gases are cooled to room temperature, and a balloon is used to maintain the constant pressure of one ATM, as is done in the chart. Once equilibrium has been achieved, calcium carbonate mixture and calcium oxide is heated to a higher temperature. How and why does the volume of the gas balloon change? The position of the equilibrium location does not change, so the storage space does not change. C The position of equilibrium moves to the left. increasing the volume. D The position of equilibrium moves to the right, reducing the volume. E The position of equilibrium moves to the right, increasing the volume. Once the balance is achieved, the total gas pressure increases to 2 ATMs. How and why does the solid mass in the heated test tubes change? The position of the equilibrium changes to the left, increasing the solid mass. BThe equilibrium position of the equilibrium does not change, so the solid block does not change. Once the equilibrium has reached its equilibrium, a small mass of calcium carbonate is added to the heated sequecon. How and why does the concentration. B Position of Equilibrium Shifts The left, increasing concentration. C The position of equilibrium moves to the right, increasing the concentration. D The position of equilibrium does not change, so the concentration. Q3: Pentachloride gas phosphorus (PCI)5 decomposes into phosphorus and trinchloride (PCI)3 and

chlorine gases, As shown in the equation: PCIPCI +CI532 The enthalpy change to this response is +87.9 kJ/mol. Does the B C D E diagram represent the rates of reactions back and forth when the pressure drops at a constant temperature? B C D E Q4: Colorless nitrogen bi-nitrogen gas tetroxide (no)24 in setted exponses in the syring is higher. B The syringe is higher. B The syringe is higher. B The syringe is higher. The syringe contains a greater number of gas molecules. The syringe contains a greaten soft gas molecules. The syringe is higher. B the syringe is higher. B the syringe is higher. The syringe is higher as on the syning is higher and allows the reaction to reach equilibrium is achieved? B C D E Qs: Chlorometan reacts with sodium iodide to produce iodine and sodium chloride, as on the other hand be in all to solubilities of sodium cloride and sodium individe in five organic solvents at 20 c. Care shown in the cale. NACI Soluble Solvents (g)M(g)Dimthyformamide 68: 100/gVO 96Formamide 62: Under certain conditions, nitric oxide (no) and nitrogen dioxide (no)2 gases react irreversibly to produce liquid hyme oxide No()+NO()ggl223 The mental change to this reaction is 41 kJ/mol. Below the boiling point of dij nitrogen troxice, which form the collowing graphs properly illustrates the effects of temperature equilibrium, assuming the thore system may reach equilibrium. Assuming the temperature of the gas

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