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## Molar mass of sodium hydrogen sulfate

Sodium sulfate name: Sodium sulfate Example reaction:  $\bullet \text{NaHSO}_4 = \text{NaOH} + \text{SO}_3 \bullet \text{NaOH} + \text{H}_2\text{SO}_4 = \text{NaHSO}_4 + \text{H}_2\text{O} \bullet \text{H}_2\text{S}_4 + \text{NaCl} = \text{HCl} + \text{NaHSO}_4 \bullet \text{H}_2\text{SO}_4 + \text{NaHCO}_3 = \text{CO}_2 + \text{NaHSO}_4 + \text{H}_2\text{O} \bullet \text{NaNO}_3 + \text{H}_2\text{SO}_4 = \text{NaHSO}_4 + \text{HNO}_3 \bullet \text{NaI} + \text{H}_2\text{SO}_4 = \text{HI} + \text{NaHSO}_4 \bullet \text{H}_2\text{O} + \text{Br}_2 + \text{NaHSO}_3 = \text{NaHSO}_4 + 2\text{HBr} \bullet \text{NaNO}_3 + \text{NaHSO}_4 = \text{Na}_2\text{SO}_4 + \text{HNO}_3 \bullet \text{H}_2\text{SO}_4 + \text{NaHCO}_3 = \text{NaHSO}_4 + \text{H}_2\text{CO}_3 \bullet \text{NaClO} + \text{NaHSO}_3 = \text{NaCl} + \text{NaHSO}_4$  Not to be confused with sodium bisulfite. Sodium bisulfate Name IUPAC Name Sodium Sulfate Bisulfate of IDENTIFIERS CAS number 7681-38-1 Y10034-88-5 (monohydrate) Y 3D model (JSmol) Interactive image ChemSpider 56397 Y ECHA InfoCard 100.028.787 EC Number 231-665-7 E number E514(ii) (acidity regulators, ...) PubChem CID 516919 RTECS number VZ1860000 UNION BU8V88OWIQ Y3KLS9Y77YJ (monohydrate) Y CompTox Dashboard (EPA) DTXS ID3033983 InChI InChI=1S/Na.H2O4S/c:1-5(2,3)4/h:(H2,1,2,3,4 )/q+1;/p-1 Key: WBHQBSYUJJSRZ-UHFFFAOYSA-M YInChI=1/Na.H2O4S/c:1-5(2,3)4/h:(H2,1,2,3,4)/q+1;/p-1Key: WBHQBSYUJJSRZ-REWHXWFOFAO SMILES [Na+]. It's all right. S(=O)(=O)O Properties Chemical formula NaHSO4 Molar mass 120.06 g/mol (anhydrous)138.07 g/mol (monohydrate) Appearance white fixed density 2,742 g/cm3 (anhydrous) 1,0 8 g/cm3 (monohydrate) 58,5 °C (331,6 K) (monohydrate) 315 °C (anhydrous) The boiling point extends to Na2S2O7 (+ H2O) at 315 °C (599 °F; 588 K) Water solubleness 28.5 g/100 ml (25 °C) 100 g/100 ml (100 °C) Solubleness acidity ammonia (pKa) 1.99 Structure Crystal structure triclin (anhydrous) monoclin (monohydrate) Hazard Safety datasheet External MSDS R-phrases (obsolete ) R34 R37 R41 S-phrases (obsolete) S26 S36 S37 S39 S45 NFPA 704 (fire diamond) 0 2 1 Flash point Non-combustible Related compounds Other sodium sulfate anions Other cations draum bisulfate are given data on materials in standard condition (at 25 °C [77 °F] , 100 kPa). Y verify (what is YN?) Infobox refers to sodium bisulfate, also known as sodium sulfate,[a] is a sodium salt of bisulfate with the molecular formula NaHSO4. Sodium bisulfate is an acid salt formed by partial neutralization of sulphuric acid equivalent to sodium base, usually in the form of sodium hydroxide (lyme) or sodium chloride (table salt). It is a dry granular product that can be safely transported and stored. The anhydrous form is hygroscopic. Sodium bisulfate solutions are acidic, with a solution of 1M with a pH of around 1. The production of sodium bisulfate is produced as an intermediate in the Mannheim process, an industrial process involving the reaction of sodium chloride and sulphuric acid:[1]  $\text{NaCl} + \text{H}_2\text{SO}_4 \rightarrow \text{HCl} + \text{NaHSO}_4$  This step is highly exothermic. Liquid sodium bisulfate is sprayed and cooled to form a solid beeper. hydrogen chloride gas is dissolved in water to produce hydrochloric acid as a useful by-product of the reaction. Even if it's not of business interest, sodium bisulfate can be generated as  $\rightarrow$  by-product of the production of many other mineral acids through the reaction of their sodium salts with excess sulphuric acid:  $\text{NaX} + \text{H}_2\text{SO}_4 \rightarrow \text{NaHSO}_4 + \text{HX}$  (  $\text{X}^- = \text{CN}^-$ ,  $\text{NO}_3^-$ ,  $\text{ClO}_4^-$ ) Produced HX acids have a lower boiling point than the reactants and are separated from the reaction mixture by distillation. Chemical reaction Hydrated sodium bisulfate dehydrates at 58 °C (136 °F), and at this time it separates from the water molecule attached to it. After cooling again, it is freshly hygroscopic. Heating sodium bisulfate to 280 °C (536 °F) produces sodium pyrosulfate, another colourless salt:[1]  $2 \text{NaHSO}_4 \rightarrow \text{Na}_2\text{S}_2\text{O}_7 + \text{H}_2\text{O}$  The use of sodium bisulfate is mainly used to reduce pH. It is also used in metal treatments, cleaning products,[2] and to reduce the pH of water for effective chlorination in swimming pools and hot tubs. Sodium bisulfate is also approved as an additive for general use, including pet food. It is used as a donkey, which urine to reduce urinary stones in cats. It is highly toxic to some embarex, but quite harmless to most other life forms; it is used in the control of starfish outbreaks. Sodium bisulfate was the main active ingredient in toilet bowl cleaners and Vanish and Sani-Flush, both of which have now stopped. [3] In the textile industry, it is sometimes applied to a velvet fabric made of silk pad and a pile of cellulose-based fibres (silk, cotton, hemp, etc.) to create burnout of velvet: sodium bisulfate, when applied to such a substance and heated, causes cellulose-based fibers to become brittle and flaked, leaving burned areas in the finished material, usually in attractive samples. [4] Sodium bisulfate is the active substance in some granular poultry bedding used to control ammonia. [5] Sodium bisulfate has also been shown to significantly reduce the concentration of Campylobacter and Salmonella in chicken domes. [6] Sodium bisulfate is used in foodstuffs as a food additive for mixtures of scamth cakes (to increase)[7] and is used in the processing of meat and poultry and, most recently, in preventing browning of fresh cut products. [8] Sodium bisulfate is considered safe by the FDA (GRAS). The food product also complies with the requirements set out in the Chemicals Code. It is marked e number E E514ii in the EU and is also approved for use in Australia, New Zealand, Canada and Mexico. [9] if it is listed as additive 514. The food class of sodium bisulfate is used in various food products, including beverages, dressings, sauces and fillings. It has many synonyms, including sodium bisulfate, sodium sulfate, sodium sulfate, hydrogen sulfate, sodium hydrosulfate and sodium sulphuric acid salt (1:1). Sodium bisulfate is considered a natural association of food manufacturers (GMA) because it is made of minerals. [11] [non-primary source needed] However, all commercially available sodium bisulfate is produced from sulphuric acid synthesized from elemental sulphal acid by the contact process. Sodium bisulfate reduces acidic-free pH and is used in place of citric,

apple or phosphoric acids, which are commercially available and can also be used as anti-browning. [13] Notes ^ The prefix bi in bisulfate comes from an outdated naming system and is based on the observation that there are twice as many sulphate soran (SO4) and other sodium sulfate as sodium sulfate (Na2SO4) and other sulphates in sodium bisulate (NaHSO4). Reference ^ and b Helms Plessen (2000). Sodium sulfate. Ullmann's encyclopedia of industrial chemistry. Weinheim: Wiley-VCH. doi:10.1002/14356007.a24\_355. ISBN 978-3527306732. † John Toedt, Darrell Koza, Kathleen Van Cleef-Toedt Chemical composition of everyday products p. 147 ^ SANI-FLUSH® Powder (discontinued), Reckitt Benckiser. ^ Margo Singer (July 11, 2007). Decoration of textile surface: Silk and velvet. University of Pennsylvania. p. 35. ISBN 978-0-8122-2000-1. † Blake, John P. Treatment of garbage for poultry (PDF). aces.edu Alabama Cooperative Expansion System. ^ Campylobacter and Salmonella Populations Associated with Chickens Raised on Acidified Litter . ^ GRAS Notice 000003: Sodium Bisulfate - FDA (PDF). Fda. ^ Sodium bisulfate - USDA (PDF). Usda. ^ Australia New Zealand Food Standards Code - Standard 1.2.4 - Labelling ingredients. www.legislation.gov.au. ^ Wise eating, easy. Noshly what flies to noshly : GMA response to natural sodium bisifla (PDF). ^ Water acidification (PDF). ^ Ali, Hussein M.; El-Gizawy, Ahmed M.; El-Bassiouny, Rawia E. I.; Saleh, Mahmoud A. (2017-05-04). Mechanisms of inhibition of browning with cysteine, asorbic acid and citric acid and identification of reaction preparations PPO-catechol-cysteine. Journal of Food Science and Technology. 52 (6): 3651-3659. doi:10.1007/s13197-014-1437-0. ISSN 0022-1155. PMC 4444905. PMID 26028748. External references Food Chemicals Codex Derived from molar matter NaHSO4 = 120.06031 g/mol This compound is also known as sodium sulfate or sodium bisulfate. Convert grams of NaHSO4 to moles or moles NaHSO4 to grams Molecular weight calculation: 22.98977 + 1.00794 + 32.065 + 15.9994 \*4 » Percent composition according to element symbol Atomic mass # atoms weight percentage sodium At 22.98977 0 1 19.149% Hydrogen H 1.00794 1 0.840% Oxygen O 15.9994 4 53.305% Sira S 32.065 1 26.707% » Similar chemical formulas Note that all formulas are case sensitive. You wanted to find weight of one of these similar formulas? NaHSO4 NaHsO4 In chemistry, the mass of a formula calculated by multiplying the atomic mass (in atomic mass units) of each element in the chemical formula by the number of atoms of that element present in the formula and then adding all of these products together. The search for molar matter begins with units of grams per mole (g/mole). When calculating the molecular weight of a chemical compound, it tells us how many grams there are in one mole of this substance. The weight of a formula is simply the mass of all atoms in a given formula in units of atoms. The atomic mass used on this site comes from NIST, the National Institute of Standards and Technology. We use the most common isotopes. This is how to calculate molar mass (average molecular weight), which is based on isotropically weighted diameters. This is not the same as molecular weight, which is the mass of one molecule of well-defined isotopes. For mass schiometric calculations, we usually determine the molar mass, which can also be called the standard atomic mass or the average atomic mass. The weights of the formula are especially useful in determining the relative weight of reagents and products in a chemical reaction. These relative weights calculated from the chemical equation are sometimes called equation scales. Using the chemical formula of the compound and the periodic table of elements, we can add up the atomic mass and calculate the molecular weight of the substance. If the formula used in the calculation of molar matter is a molecular formula, the calculated weight of the formula is the molecular weight. The percentage by mass of any atom or group of atoms in a compound can be calculated by dividing the total mass of the atom (or group of atoms) in the formula weight formula by multiplying it by 100. A common requirement on this page is to convert grams into moles. To complete this calculation, you need to know what substance you are trying to convert. This is because the molar weight of the substance affects conversion. This page explains how to find molar matter. Mass.

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