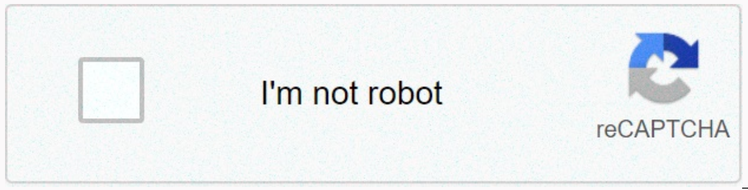

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Acid base test kit 1 answers

Einige Word-Funktionen können in Google Docs nicht angezeigt werden und werden bei Änderungen entferntDetails anzeigenLetzte Änderungen anzeigen In order to continue enjoying our site, we ask you to confirm your identity as a human being. Thank you very much for your cooperation. Solved the best collection of acids, bases and salts class 10 Science Important Questions and Answers Chapter 2 Pdf from the latest books edition of NCERT, this will help you in scoring more grades in CBSE Exams. Class 10 Science Chapter 2 Important Questions with Acid Answers, Basics and Salt Class 10 Chemistry Chapter 2 Important Questions With Acid Answers, Bases and Acid Salt, Basics and Salt Class 10 Important Questions Very Short Answer Type Question 1. How does hydronic ion concentration (H3O+) affect when acid solution is diluted? (2017 OD) Answer: When the acid solution is diluted with water, the concentration of ions (H3O+) decreases. Question 2. Why does HCl dry gas not change the color of dry litmus paper? A: Since HCl gas gives H+ ions only with H2O molecules behave like acid. Dry HCl does not change the color of dry litmus paper because it needs moisture or water for its sour action. Two A and B solutions have pH values of 5 and 8 respectively. What solution will be basic in nature? Answer: Solution B will be the main one by nature, as its pH value is 8. Question 4. Fresh milk has pH 6. When it changes to cheese (yogurt), will its pH value increase or decrease? Why? Answer: When milk changes to cheese, its pH values will decrease as the cheese contains lactic acid (pH < 6 acids). Acids, Foundations and Salt Class 10 Important Questions Short Answer Type I Question 1. The state has the number of water molecules present in the crystals of washing soda and plaster of Paris. What are these water molecules called? (2013) Answer: Detergent soda crystals contain 10 water molecules (Na2CO3 . 10H2O). The plaster crystals of Paris contains 1/2 water molecules (CaSO4.1/2 H2O or 2CaSO4 . H2O). Water molecules that form part of the crystal structure are called crystallization water. Question 2. Two A and B solutions have pH 3 and 5 respectively. Which of the two solutions has a greater concentration of hydrogen ions, and which is more acidic? Give rise to your answer. (2013) Response: pH solution A = 3, pH solution B = 5 Solution A is more acidic because solution A has a greater concentration of hydrogen ions, since the pH of the solution is inversely proportional to the concentration of hydrogen ions in it. That is, the solution, which has a low concentration of hydrogen ions, has a higher pH value. Question 3. Explain why the water solution for sodium sulfate is neutral, while the water solution of sodium carbonate is basic in nature. (2014) Answer: Aqueous solution of sodium sulfate gets hydrolyzed to form sodium hydroxide and acids, which are a strong base and strong acid respectively. So its water solution is neutral. When sodium carbonate dissolves in water, it hydrolyzes to some extent and forms sodium hydroxide and carbon dioxide. Now sodium hydroxide is a solid base that is fully ionized and gives a large number of hydroxide ions [OH– (aq)]. On the other hand, carbon dioxide is a weak acid that is only slightly ionized and therefore gives a small amount of hydrogen ions [H+ (aq)]. The mixture contains more hydroxide ions than hydrogen ions, so is basic in nature. Question 4. Name the gas that is usually produced when the dil. Sulfuric acid reacts with metal. Illustrate it by example. How will you test the evolution of this gas? (2015) Answer: Metals react with a deed of sulfuric acid to give metal sulfates and hydrogen gas. For example, Zn (s) + H2SO4 (aq) → ZnSO4 (aq) + H2 ↑ (Hydrogen gas) Hydrogen gas test: Take about 5 ml of diluted sulfuric acid in a test tube and add a few pieces of zinc pellets to it. Hydrogen gas evolves, which forms bubbles in soap solution. Bring a burning candle near a hydrogen-filled bubble. It burns with pop sound. Question 5. Write a chemical equation to describe how diet soda is produced on a large scale. Also write the chemical name of the products obtained. (2014) Answer: The chemical name of diet soda is sodium bicarbonate (NaHCO3) or sodium hydrogen carbonate. Sodium hydrogen carbonate is produced on a large scale, reacting to a cold and concentrated solution of sodium chloride (called brine) with ammonia and carbon dioxide. Acids, Bases and Salt Class of 10 Important Questions Short Answer Type II Question 1. (a) Write a name given to the basics that are very soluble in water? Give an example. (2012, 2014) (b) How is tooth decay related to pH? How can this be prevented? (c) Why does stinging beat cause pain and irritation? Rubbing baking soda on the area of stinging gives relief. As? Answer: (a) Water-solub bases are called alkale, for example, NaOH, KOH. (b) Tooth decay begins when the pH of the mouth is below 5.5. Bacteria present in the mouth produce acids by degrading sugar and food particles left in the mouth after eating. This can be prevented by cleaning the mouth and teeth with a tooth pass, which are usually basic and can neutralize excess acid. (c) When a honeybee absorbs a person, it injects acidic fluid into the skin, which causes tremendous pain and irritation. Rubbing a soft base like a solution of baking soda on the tungsten skin area gives relief. This is because, being the basis, diet soda neutralizes the acidic liquid administered by the butnis and reverses its effect. Question 2. (a) You are given two solutions A and B. RN solution A is 6, and pH solution B - 8. (i) Identify an acidic and basic solution. (ii) Which solution has more concentration of H+ Ions? Give for your answer. (b) Why is HCl a stronger acid than acetic acid? Explain. (2012) Answer: (a) Solution A: pH = 6 Solution B: pH = 8 (i) A is an acidic solution, while B is the main solution. (ii) Solution A has a higher concentration of H+ Ions, as the higher concentration of hydronium ions is lower than the pH value. (b) HCl is a stronger acid than acetic acid because acid strength depends on the amount of H+ ions produced. When the concentration of HCl and acetic acid is taken, the two produce different amounts of hydrogen ions. HCl produces more H+ ions and is said to be a strong acid while acetic acid yields fewer H+ions, so said to be a weak acid. Question 3. Reason for condition: (2014) (i) dry gas HCl does not change the color of dry blue litmus paper. (ii) alcohol and glucose also contain hydrogen, but do not conduct electricity. (iii) Cone, H3O+ ion affects the dilution of acid solution. Answer: (i) HCl dry gas does not contain hydrogen ions in it, so it shows no acidic behavior. HCl dry gas does not change the color of dry blue litmus paper because it does not have hydrogen ions [H+ (aq)]. However, when HCl gas dissolves in water, it forms hydrogen ions and therefore shows acidic behavior. (ii) Aqueous glucose and alcohol solutions do not show an acidic nature, because their hydrogen is not separated as hydrogen ions [H+ (aq) ions] when dissolved in water. So, they don't conduct electricity. (iii) Acid is a substance that is dissocified when dissolved in water to produce hydrogen ions [H+ (aq) ions], hydrogen ions do not exist as H+ ions in the solution, they are attached to polar water molecules to form hydronium ions, H3O+. H+ + H2O → H3O+ When the acid solution is diluted, the number of ions [H3O+] per unit decreases. Thus, the pH of the solution increases. Question 4. Illustrate any three chemical properties of acids. With examples. (2014) Answer: Chemical properties of acids: Acids react with metals to form H2 gas. Zn + 2HCl → ZnCl2 + H2 ^ Acids turn blue litmus into red. Acids react with the basics to the formation of salt and water (neutralization reactions). Question 5. What three chemicals are obtained when conducting electricity through an aqueous brine solution? Write one industrial use each. (2015) Answer: When electricity passes through a concentrated solution of sodium chloride (called brine), it decomposes to form sodium hydroxide, chlorine and hydrogen gas. Use of sodium hydroxide (NaOH). Used for cooking soap and detergents. Chlorine use (Cl2). It is used in the production of bleaching powder. Use of hydrogen gas (H2). Used for application of ammonia for fertilizers; used as fuel or margarine. Question 6. When soap is scrub on a slick of curry on a white cloth, why it turns reddish-brown, and turns yellow again when the fabric is with a lot of water? (2015) Answer: The curry contains turmuma, which acts as an acid base indicator. Turoum contains yellow dye, which turns red in the main solutions when soap is cleaned on a curry stain on a white cloth. This is due to the fact that the soap solution has a basic character, which changes the color of turman in a stain of curry to reddish-brown. This stain turns yellow again when the fabric is washed with plenty of water, because then the main soap is removed by water. Question 7. You are given three tubes C, A and B, which contain distilled water, acidic solution and the main solution. If you are given only blue litmus paper, how will you identify the contents of each test tube? (2015) Answer: First, take three strips of blue litmus paper and plunged one into each test tube. The liquid in the tube turns the blue litmus into red. The test tube A contains acid because (acid turns blue litmus into red). If the other two tubes do not change the color of the blue litmus paper, it shows that one of them contains a base and the other contains distilled water. Now put the blue litmus paper, which has flushed acid test tubeS A into the other two tubes. If test tube liquid B converts this red litmus paper back into blue, it shows that it is the basis. The liquid in test tube C does not flip the color of either blue litmus or red litmus, thus distilled water. Question 8. Write the chemical name Plaster of Paris. Write the chemical equation of its preparation. Why should the plaster of Paris be stored in a dry place? (2017 D) Answer: The chemical name of The Plaster of Paris is calcium sulfate hemihydrate, and its formula is CaSO4. 1/4 H2O. Preparation: Acids, bases and salt class 10 Important questions long answer type question 1. (a) Write the chemical name and chemical formula of soda washing. (b) How is it obtained from sodium chloride? Give the equation reactions. (c) Why is it called the main salt? Give it any use. (2012) Answer: (a) Chemical Name of Detergent Soda: Sodium Carbonate Chemical Soda Wash Formula: Na2CO3.10H2O (b) Detergents produced by Solvay process by reaction of amonic brine with CO2 NaCl + NH3 + CO2 + H2O → <8> NaHCO3 + NH4Cl NaHCO3 on heating produces Na2CO3 (gum soda), which on recrystilation produces detergent soda (c) Soda washing is the main salt because it is strong base salt (NaOH) and weak acid (H2CO3) with a pH value of more than 7. Usage: Detergents are used to soften hard water. Question 2. Write the chemical name Na2CO3.10H2O and Na2CO3. Write a value of 10H2O. Remember the term used for water molecules attached with salt. Using the chemical equation, explain how to cook both Na2CO3.10H2O and Na2CO3. It will also list two uses of Na2CO3.10H2O. (2012) Answer: Chemical Name Na2CO3.10H2O: Sodium Carbonate Detergent sodas. Na2CO3: Sodium carbonate (soda) 10H2O is the water of crystallization of sodium carbonate. Water crystallization is a fixed amount of water molecules present in one formulaic unit of salt. Thus, water molecules 10 are present in one formulaic unit of sodium carbonate. Soda washing is produced by solvay process, by reaction of amonic brine with CO2 gas. NaHCO3 on heating produces Na2CO3 (sodium carbonate, called soda ash), which on recrystilization produces detergent sodas. Washing soda is the main salt. Some important applications of detergent soda: Detergents are used in the glass, soap and paper industries. It is used in the manufacture of borax compound. Used as a cleanser for domestic purposes. Used to remove constant hardness of water. Question 3. (a) Identify the calcium compound, which is a yellowish white powder and is used to disinfect drinking water. Write its chemical name and formula. How is it manufactured? Write a chemical equation for the reaction involved. Also list two other connection usages. (b) Write a balanced chemical equation of the chlorine meadows process. (2012) Answer: (a) Connecting bleaching powder. Chemical name: Calcium ocychloride Formula: CaOCl2 Produced by chlorine action on a dry slap of lime Ca(OH)2. Ca(OH)2 + Cl2 → CaOCl2 + H2O Use: It is used for whitening wood cellulose in paper factories. It is used as an edifying agent in many chemical industries. (b) When electricity passes through an aqueous solution of sodium chloride (called brine), it decomposes into sodium hydroxide formation. This process is called chlorine-alkaly process. 2NaCl(aq) + 2H2O (l) → 2NaOH (aq) + Cl2 (g) + H2 (g) Question 4. (a) Learn the following chemical equation: Name the reactive agent and product and are about one use of the product. (b) The following salts are formed by acid reaction with the base: (i) Sodium chloride (ii) Ammonium Nitrate Identify acid and base and subtab your response in the format below: S.No. Hydrochloric acid Basic nature 1. Sodium chloride 2. Ammonium Nitrate Answer: (a) Reactant: Gypsum Product: Plaster Paris and Water Use: The Plaster of Paris is used in hospitals to install bone fractures in the right position to ensure proper healing. (b) S.No. Hydrochloric acid Basic nature 1. Sodium Chloride Hydrochloric Acid (HCl) Sodium Hydroxide (NaOH) Neutral 2. Ammonium Nitrate Nitric Acid (HNO3) Ammonium Hydroxide (NH4OH) Acidic Question 5. (a) The student dropped several pieces of marble into the diluted hydrochloric acid contained in the test tube. The evolutionary gas passed through lime water. What changes will be observed in limewater? Write balanced chemical equations for both changes observed. (b) Indicate the chemical property in each case on which the the use of baking soda: (i) as an antakyd (ii) as part of the baking powder. (2014) (2014) (a) When a piece of marble (CaCO3) has fallen into dil HCl, CO2 gas will evolve, which turns lime water milky. The formation of white ppt CaCO3 makes lime water milky. (b) (i) Being alkaise, sodium hydrogen carbonate (diet soda) neutralizes the excess acid present in the stomach and relieves indistinguishment. (ii) Baking powder is a mixture of baking soda and mild edict acid such as tartaric acid. When baking powder mixes with water (present in dough made for baking cake or bread), sodium hydrogen carbonate reacts with tartaric acid to develop CO2 gas. Question 6. (a) Define a universal indicator. Mention of its one use. (2015) (b) Solution A gives a pink color when a drop of the phenolftalin indicator is added to it. Solution B gives red color when added to it a drop of methyl orange. What type of solutions is A and B and which of the solutions A and B will have a higher pH cost? (c) Name one salt, the solution of which has pH more than 7 and one salt, the solution of which has pH less than 7. Answer: (a) Universal indicator is a mixture of many different indicators, which gives different colors at different pH values of the entire pH scale. It shows different colors at different concentrations of hydrogen ions in the solution. (b) Solution A gives pink color when the phenolftalin indicator is added, so A is the basis. Solution B gives a red color when it is added a drop of methyl orange, so B is acid. Thus, solution A will have a lower concentration of hydrogen ion than B. Thus, A will have a pH of more than 7, because pH values of acid solution < 7; base solution > 7; and neutral solution = 7. (c) Salts of strong acids and weak bases give an acidic solution that has a pH of less than 7. For example, NH4Cl, ammonium chloride will have pH of less than 7. Salts of weak acids and durable bases give the main solution having pH more than 7. For example, Na2CO3, sodium carbonate will have pH of more than 7. Question 7. (a) The blue color of the substance crystals changed when heated in a closed tube, but the color was restored after once cooling. Name the substance and write its chemical formula. Explain the phenomenon involved. (b) Write the name and chemical formula of two such compounds, whose one unit of formula is associated with water molecules 10 and 2 respectively. (2015) Answer: Substance copper sulphate crystals (CuSO4.5H2O) that are blue. When copper sulphate crystals are heavily heated, they lose all crystallization water and form an anhydrotic copper sulfate (which is white): Thus, when heated strongly, blue copper sulphate crystals turn white due to water loss crystallization. Dehydration of copper sulphate crystals is a reversible process. Therefore, when water is added to the anhydrant copper sulfate, it moisturizes and turns blue due to the formation of hydrated copper sulfates. (b) Chemical formula such compounds, whose one unit of formula is associated with water molecules 10 and 2 respectively Question 8. Write an exercise to show acid reaction with metal carbonates and metallic hydrogen carbonate salts. (2017D) Answer: Take two test tubes. Mark them as A and B. Take about 0.5 um of sodium carbonate in the test tube A and 0.5 g of hydrogen carbonate in the test tube B. Add about 2ml of diluted HCl acid to both of these tubes. Pass the gas produced on a case-by-case basis through lime water, as shown in the figure. Lime water in each case will become milk due to the production of CO2 gas in each of the test tubes. Reactions occurring in the above activities: Test tube 'A' Na2CO3 (s) + 2HCl (aq) → 2NaCl (aq) + CO2 (g) + H2O (l) Test tube 'B' NaHCO3 (s) + HCl (aq) → NaCl (aq) + CO2 (g) + H2O (l) When passing gas evolved through lime water: But when passing excess CO2 gas through lime water, lacticity disappears due to the formation of soluble CaHCO3 This activity shows that metal carbonates and metallic hydrogen carbonates react with acids to salt formation , CO2 and water. Question 9. (a) What is the importance of pH in everyday life? (2017 OD) (b) How are sodium hydroxide and Cl2 (chlorine) gas produced from joint salt. What is the name of this process? Answer: (a) (i) Living organisms can only survive in a narrow range of pH changes. Acidic rainwater, when fleeing into the river, reduces the value of river water pH and complicates the survival of watercarian life in such river water. Plants require a specific pH range for their healthy growth. (ii) Our stomach and intestines work in a certain range of pH. The stomach acts in a slightly acidic environment, and the small intestine digests food in a slightly alkaly environment. (iii) Tooth decay begins when the pH of the mouth is below 5.5. (b) When electricity passes through an aqueous solution of sodium chloride (called brine), it decomposes into sodium hydroxide formation. This process is called chlorine-lugali process because the products formed by chlorine and naoh meadows. 2NaCl (aq) + 2H2O (l) → 2NaOH (aq) + Cl2(g) + H2 (g) The Cl2 gas is given to the anody and the H2 gas on the cathody while a solution of sodium hydroxide is formed near the cathody. Cathode.

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