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Chapter 17 thermochemistry test b answers

Java games: Flashcards, matching, concentration and word search. Use these activities to learn the vocabulary and the key concepts presented in this chapter. Ab calorie amount of heat that increases the temperature of 1 g of clean water by 1 ° C joulethe SI unit of energy exothermica process that loses heat to the environment molar heat of the solution heat absorbed or released at the solution of 1 mole of substance molar heat of the fusion heat is absorbed by 1 mole of a substance in melting from a solid to a liquid at a constant temperature calorie meter device used to measure the amount of heat absorbed or released during chemical or physical processes enthalpythe heat content of a substance energy capacity to do work or to deliver heat Hess law on heat summary two or more thermochemical equations provides a final equation, their heat changes can be added for the final heat change endothermic reaction chemical change where heat is absorbed the amount of heat required to change the temperature of an object by exactly 1°C law on the preservation of energyIn any chemical or physical process, energy is neither created nor destroyed reaction heat is absorbed or released by a chemical reaction specific amount of heat required to increase the temperature of 1 gram of a substance 1 °C heat energy that always flows from a warmer object to a cooler object standard heat of the formation change in enthalpy that accompanies the formation of 1 mole of a compound from its elements, with all substances in their standard states at 25 ° C calorie-proofing the exact measurement of the heat flow into or out of a system of chemical and physical processes chemical potential energy stored in chemical bonds of a substance heat of combustion heating of reaction (enthalpy change) for complete burning of a mole of a substance molar heat of condensation heat released when a mole of steam condenses at the normal boiling point molar heat of solidification heat lost when a mole of a liquid hardens at a constant temperature molar heat of the amount of heat needed to evaporate a mole of a liquid environment no matter what happens in the universe minus the system that focuses on thermochemical equation chemical equation that includes enthalpy change systemthe part of the universe as a scientist focuses his attention thermochemistry study of energy changes occurring during chemical reactions and changes in condition

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