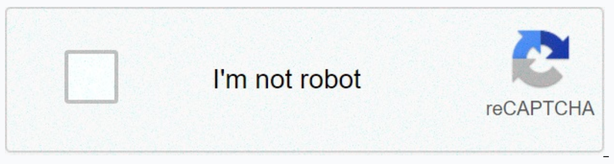

Continue

Iron ii oxide negative ion

1) ZnSO4 = cinko sulfatas 2) Al2(SO4)3 = aliuminio sulfatas 3) Ag2Cr2O7 = sidabro dichromatas 4) FeSO4 = geležis(II)sulfatas 5) PbSO4 = švino sulfatas 6) Fe2(SO4)3 = geležies (III)sulfatas 7) K2CO3 = kalio karbonatas 8) Cu(OH)2 = vario hidroksidas 9) Na2CO3 = natrio karbonatas 10) Al2S3 = aliuminio sulfidas 11) NaHS = natrio vandenilio sulfidas 12) Fe2(SO3)3 = geležies (III) sulfitas Spustelėkite, jei norite paslėpti tirpalus Fe2(CO3)3 (NH4)2CO3 Li2CO3 Ca3PO4 Fe(OH)2 Fe(OH)3 NiSO4 NiCl2 Ag2CO3 Ag2S (NH4)3N PbCO3 Al(OH)3 Spustelėkite, jei norite paslėpti Monoatomic Ions fiksuoto įkrovimo simbolio simbolio pavadinimo pavadinimas H+ vandenilio jonų H- hidrido ličio jonų F- fluorido Na+ natrio jonų cl- chlorido K+ kalio jonų Br- bromidas Rb+ rubidžio jonų I- jodido Cs+ cezio jonų O2- oksidas Be2+ berilio jonų S2- sulfidas Mg2+ magnio jonų Se2- selenidas Ca2+ kalcio jonų Te2- katrincidas Sr2+ stroncio jonų N3- nitride Ba2+ bario jonų P3- fosfidas Ra2+ radžioionas As3- arsenic Ag+ silverion Zn2+ zinc Al3+ aluminium Note that the letters in the name of ion before the -ide limb are the stem. For example, the bromide stem is bromine. Moreover, only in case phosphide P is a major part of P. Symbols and fees for Monoatomic Ions Variable Charge System name Common systemic name Common symbol (Stock system) name Symbol (Stock system) name Cu+ copper(I) cuprous Hg22+ mercury(I) mercury Cu2+ copper (II) copper (II) copper H g2+ mercury(II) mercury Fe2+ iron(II) black Pb2+ lead(II) plumbing Fe3+ iron(III) iron Pb4+ lead (IV) plumbing Sn2+ stannous Co2+ cobalt(II) cobalt Sn4+ tin(IV) stannic Co3+ cobalt(III) cobalt Cr2+ chromium(II) chromium Ni2+ nickel(II) nickel Cr3+ chromium(III) chromium Ni4+ nickel(IV) nickel Mn2+ manganese(II) manganese Au+ gold(I) aurous Mn3+ manganese(III)) manganese Au3+ gold(III) aura symbols and charges for polyatomic ions formula name Formula name NO3-nitrate ClO4-perchlorate NO2-nitrite ClO3-chlorate CrO42 - chromate clo2-chlorite Cr2O72-dichromate clo-hypochlorite CN-cyanide IO4-periodate MnO4-permanganate IO3-iodate OH- hydroxide IO-hypothydite O22-peroxide BrO3-bromate NH2-amide BrO-hypobromite CO32-carbonate HCO3-hydrogen carbonate (bicarbonate) SO42-sulphate HSO4 hydrogen sulphate (bisulphate) SO32-sulphite HSO3- hydrogen sulphite (bisulphite) C2O42-oxalate HC2O4 - hydrogen oxalate (binoxalate) PO43 Phosphate HPO42-hydrogen phosphate PO33-phosphite H2PO4-dihydrogen phosphate S2O32-thiosulphate HS-hydrogen sulphide AsO43-arsenic BO33-bor SeO42-sele-selenium B4O72-tetraborate SiO32-silicate SiF62-hexafluoroosilic ate C4H4O62-tartrate C2H3O2-acetate (an the way to write acetate is CH3COO-) There is one positive polyatomic. It is NH4+ and is called ammonium ion. Note: Only a plus sign or minus sign for ions with +1 or -1 fees is acceptable. Binary compounds of metals with fixed fees Given formula, write name binary compound is one of two different elements. There may be one of each item, such as NaCl or KF. There may also be several items for each item, such as Na2O or AlBr3. Please note that all items participating in this lesson have only one charge. This includes both formula and anion. Specifies to remember about the name of a compound from its formula 1. The binary compound in the order of the names is first a vocation, then anion. 2. Use the cation name with fixed oxidation state directly from the periodic table. 3. The name of the anion will be made from the root of the element name and the arthmeal -ide. Example 1: Write the name of this formula: H2S step #1 - Look at the first item and name it. The result of this step = hydrogen. Step #2 - Look at the second element. Use the root of your full name (which is sulf-) plus ends -ide. The result of this step = sulfide. These two steps give the full name H2S. Note that the presence of the subscript is ignored. There are other types of binary compounds where you have to pay attention to describe. These compounds include variable charges. Your teacher will tell you which ones you will be responsible for. Example 2: Write the name of this formula: NaCl step #1 - look at the first item and name it. The result of this step = sodium. Step #2 - Look at the second element. Use the roots in your full name (which is chlor-) plus ends -ide. The result of this step = chloride. Example 3: Write the name of this formula: MgBr2 step #1 - Look at the first item and name it. The result of this step = magnesium. Step #2 - Look at the second element. Use the root of your full name (which is bromo-) plus ends -ide. The result of this step = bromide. Note that the presence of a subscript in this name is not affected. Example 4: Type the name of the following formula: KCl The first part of the name is from the first character of the element: Potash. The second part of the name is obtained from the root plus -ide of the second character, so chlorine + ide = chloride. This compound is named potassium chloride example 5: Type the name of this formula: Na2S The first character is Na, so the first part of the name is sodium. (Note that the presence of a subscript in this title is not affected.) The second element is sulfur (from the symbol S), so the name is sulf + ide = sulfide. This compound is named sodium sulphide. Three possible errors to be made 1) Often students forget to use the arffix -ide. BaS, for example, is named barium barium An ignorant student may want to call it barium sulfur. 2) Make sure the second name is root plus -ide. An ignorant student may want to name BaS as barium sulfur. NaBr is not named sodium bromide, the main answer is sodium bromide. 3) There is a set of binary compounds, which is named using Roman numerals. Pupils often confuse two sets of rules. For example, a student may want to be called Na2S sodium (I) sulfide. Although it is never wrong to use Roman numerals, your teacher will probably want you to use Roman numerals only in certain cations. Here are four examples of common roots: Cl: chlor- F: fluor- Br: brom- O: ox- Practice Problems Write the correct name for: Answers 1) MgS 2) KBr 3) Ba3N2 4) Al2O3 5) NaI 6) SrF2 7) Li2S 8) RaCl2 9) CaO 10) AlP 11) K2S 12) LiBr 13) Sr3P2 14) BaCl2 15) NaBr 16) MgF2 17) Na2O 18) SrS 19) BN 20) AlN 21) Cs2O 22) RbI 23) MgO 24) CaBr2 25) LiI 26) BeBr2 27) K2O 28) SrI2 29) BF3 30) Al2S3 1) magnesium sulfide 2) potassium bromide 3) barium nitride 4) aluminum oxide 5) sodium iodide 6) strontium fluoride 7) lithium sulfide 8) radium chloride 9) calcium oxide 10) aluminum phosphide 11) potassium sulfide 12) lithium bromide 13) strontium phosphide 14) barium chloride 15) sodium bromide 16) magnesium fluoride 17) sodium oxide 18) strontium sulfide 19) boron nitride 20) aluminum nitride 21) cesium oxide 22) rubidium iodide 23) magnesium oxide 24) calcium bromide 25) lithium iodide 26) berylium bromide 27) potassium oxide 28) strontium iodide 29) boric fluoride 30) aluminium sulphide Binary metal compounds with fixed taxes Given name, write the formula example 1: Write a formula from this name: sodium bromide step #1 - Write down the symbol of the first word and charge. Result = Na+ #2 - Write down the symbol and boot of the second word. Result = Step Br #3 – Use the minimum number of callings and anions required to have the sum of all taxes in the formula to zero. In this case, only one Na+ and one Br¯br. The resulting formula is NaBr. Example 2: Write a formula from this name: potassium chloride #1 step – write down the symbol and boot of the first word. Result = K+ #2 - Write down the symbol and boot of the second word. Result = Step #3 – Use the minimum number of callings and anions required to have the sum of all taxes in the formula to zero. In this case, you only need one K+ and¯Cl. The resulting formula is KCl. Example 3: Write the formula from this name: barium iodide #1 step is to write down the character and boot of the first word. Result = Ba2+ #2 – Write down the symbol and boot of the second word. Result = Step #3 – Use the minimum number of callings and anions required to have the sum of all taxes in the formula to zero. In this case, it is necessary to one Ba2+, but two are Why? Answer - Two negative one charges are necessary because there is one postive two charge. Only in this way can the total formula tax be zero. The resulting formula is BaI2. Example 4: Write the formula from this name: aluminum chloride #1 step - write down the symbol and boot of the first word. Result = Al3+ #2 – write down the symbol and boot of the second word. Result = Step #3 – Use the minimum number of callings and anions required to have the sum of all taxes in the formula to zero. In this case, only one Al3+ is required, but you need¯Cl. why? Answer - Three negative one charges are necessary because there is one postive three charge. Only in this way can the total formula tax be zero. The resulting formula is AlCl3. Example 5: Write the name of the following formula: Magnesium oxide #1 step – Write down the symbol and boot of the first word. Result = Mg2+ #2 – write down the symbol and charge of the second word. Result = O2 step #3 – Use the minimum number of callings and anions required to have the sum of all taxes in the formula to zero. In this case, one Mg2+ and one O2 is required. Answer - One positive two charge is offset by one negative two charge. This gives a zero total formula fee. The formula resulting is MgO. Example 6: Write the name of the following formula: Alumina #1 - Write down the symbol and charge of the first word. Result = Al3+ #2 – write down the symbol and boot of the second word. Result = O2 step #3 – Use the minimum number of callings and anions required to have the sum of all taxes in the formula to zero. In this case, you need two Al3+ and three O2¯. why? The answer is that this is the only possible way to get positive and negative taxes equal and keep the numbers to a minimum. Please note that the positive fee is +6 and the negative fee is -6. Also, keep in mind that you cannot change the charges to correct the formula. The resulting formula is Al2O3. Warning: beware of the temptation to write the above formula as Al3O2. Practice problems Write the correct formula: Answers 1) magnesium oxide 2) lithium bromide 3) calcium nitride 4) aluminium sulphide 5) potassium iodide 6) strontium chloride 7) sodium sulphide 8) radium bromide 9) magnesium sulphide 10) aluminium nididide 1 (1) cezi sulfide 12) potassium chloride 13) strontium phosphide 14) barium iodide 15) sodium fluoride 16) calcium bromide 17) beryllium oxide 18) strontium sulphide 19) boric fluoride 20) aluminium phosphoid 21) rubidium oxide 22) calcium iodide 23) cezioxide 24) magnesium iodide 25) lithium chloride 26) beryllium bromide 27) sodium oxide 28) calcium fluoride 29) boron phosphide 30) alumina 1) MgO 2) LiBr 3) Ca3N2 4) Al2S3 5) KI 6) SrCl2 7) Na2S 8) RaBr2 9) MgS 10) AlN 11) Cs2S 12) KCl 13) Sr3P2 14) BaI2 NaF 16) CaBr2 17) BeO 18) SrS 19) BF3 20) AlP 21) Rb2O 22) CaI2 23) Cs2O 24) Mg LiCl 26) BeBr2 27) Na2O 28) CaF2 29) BP 30) Al2O3 binary cation compounds with variable charges The formula given, please enter a name. Stock System A binary compound is one of two different elements. There may be one of each item, such as CuCl or FeO. There may also be several items for each item, such as Fe2O3 or SnBr4. All four above formulas are examples of this type. The lessons involved in this lesson have at least two accusations. The anions involved have only one charge. The type of names you'll find out is called a stock scheme or a Stock Scheme. It was created by the German chemist Alfred Stock (1876-1946), which was first released in 1919. In his own words, he felt that the system was simple, clear, immediately understandable, capable of being the most general. 1924 The German Commission recommended the approval of the Stock scheme with certain amendments. For example, FeCl2, which according to the original Stock idea would have been named iron (2) chloride, became iron (II) chloride in the revised proposal. 1934 Stocks confirmed roman digits, but it felt better to keep the hyphen and drop in brackets. This proposal has not been complied with, but the stock system continues to be used worldwide. Example #1: Type a name: FeCl2 step #1 - the first part of the name is the unchanged name of the first formula element. In this example, it would be iron. Step #2 - after the first step, the result will be a Roman numeral. Here's how to adjust its value: 1. Multiply the anion (Cl) charge by its subscript. Ignore the fact that it is negative. In this example, it is one time two equals two. divide this result by the lower index (Fe). This is the value of using Roman numerals. In this example, it is divided by one equal to two. 3. The roman number value is equal to the positive calling charge of this formula. Since the result of the step is #2 2, we use iron (II) in the title. Note that there is no space between the name and the parenthese. Step #3 - anion named after the usual stem way plus ide. The correct name of the sample is iron (II) chloride. Example #2: Name this compound: CuCl2 In this example, I explained it differently. Compare it with the above. This #4 also explains the example of #4. · The first part of the name is from the first character of the element: copper. · Roman number is II, because 2 chlorides equal to -2, so Cu must be +2. (It must be +2 so that the total charge is zero. · The second part of the title comes from the roots of the second character ide, so chlorine + chloride = chloride. This compound is named copper (II) chloride. Example #3: Write a name: Fe2O3 #1 step - The first part of the name is the unchanged name of the first formula element. In this example, it would be iron. Step #2 - after the first step, the result will be a Roman numeral. Here's how to set its value: 1. Multiply the anion (O) charge by its subscript. Ignore the fact that it is negative. In this example, it is twice three equal to six. 2. Divide this result by the lower index (Fe). This is the value of using Roman numerals. In this example, it is six divided by two levels of three. 3. Note: This value of the Number of Romans is equal to the positive fee for the calling. In this example, the result of the step is #2 3. This means that iron (III) will be used for the name. Note that there is no space between the name and the parenthese. Step #3 - anion named after the usual stem way plus ide. The correct name of the sample is iron (III) oxide. Example #4: name this compound: SnO · The first symbol is Sn, so the first part of the name is tin. · The Roman number is II, because one oxygen = -2, so one tin equals +2. · The second element is oxygen (from the symbol O), so the name is taurus + ide = oxide. This compound is named tin (II) oxide. Practice issues Reply using the inventory system. Write the correct name for: Answers 1) NiS 2) PbBr4 3) Pb3N2 4) Fe2O3 5) FeI2 6) Sn3P4 7) Cu2S 8) SnCl2 9) HgO 10) Hg2F2 11) CuCl2 12) CuBr 13) PbO 14) Fe2S3 15) NiCl2 16) SnO 17) Cu2O 18) PbO2 19) NiO2 20) SnO2 21) Hg2O 22) Hg2I2 23) AuCl3 24) MnO 25) CrCl3 26) CoO 27) Mn2O3 28) Co2S3 29) AuF 30) CrBr2 1) nickel(II) sulfide 2) lead(IV) bromide 3) lead(II) nitride 4) iron(III) oxide 5) iron(II) iodide 6) tin(IV) phosphide 7) copper(I) sulfide 8) tin(II) chloride 9) mercury(II) oxide 10) mercury(I) fluoride 11) copper(II) chloride 12) copper(I) bromide 13) lead(II) oxide 14) iron(III) sulfide 15) nickel(II) chloride 16) tin(II) oxide 17) copper(I) oxide 18) lead(IV) oxide 19) nickel(IV) oxide 20) tin(IV) oxide 21) mercury(I) oxide 22) mercury(I) iodide 23) gold(III) chloride 24) manganese((II) oxide 25) chromium(III) chloride 26) cobalt (II) oxide 27) manganese(III) oxide 28) cobalt (III) sulphide 29) gold(I) fluoride 30) chromiuno (II) bromide There are certain formulations, which you will never reduce the most common is mercury(I), which is Hg22+. It never hg + itself. It always comes as a couple. There are reasons for this behavior, but it does not fall within the scope of this work. Example #1: Mercury(I) chloride The formula of this compound is Hg2Cl2. It is not HgCl. You will be tagged incorrectly for the test to do so. Example #2: Mercury(I) nitrate The formula for this compound is Hg2(NO3)2. Once again, that not be reduced. Why? In nature, mercury (I) is a set of two atoms, not just one. The relevant mercury (II) formula for two samples would be HgCl2 and Hg(NO3)2. The second main category is peroxide, which is O22¯. Example #3: Hydrogen Peroxide Formula is H2O2 and it does not decrease to HO. Why? The same reason as above, peroxide travels as a group of two oxygen atoms, not one. Example #4: Sodium peroxide Formula is Na2O2. For binary kation compounds with variable charges to which the name is given, type a formula. Example of the stock system #1 - Write a formula: copper (II) chloride #1 step - the first word will tell you the symbol of the calling. In this case, this is a step in Cu. #2 - the Roman numeral will tell you the fee for calling. In this case, it is a positive two. Step #3 - anion symbol and charge comes from the second name. In this case, chloride means Cl. Step #4 - remembering the rule that the formula should have a zero total charge, you write the formula CuCl2. ChemTeam often asks students: But how do you know that chloride means Cl¯? This type of question is usually answered with a question, as do you know the name and face of your best friend? Correctly, you spent time in your company to the point where the memorable relationship between name and face. Chloride is the name and Cl is the face. Example #2 - Write a formula: copper (I) oxide #1 step - the first word will tell you the calling symbol. It's Cu. step #2 - the Roman numeral will tell you about the calling fee. He is positive. Step #3 - anion symbol and charge comes from the second name. Oxide means O2¯. Step #4 – since the formula must have a zero total charge, you write the formula Cu2O. Example #3 - Write a formula: iron (III) #1 step - the symbol of the vocation is Fe. Step #2 - the vocation fee is a positive three. Remember that comes from Roman numerals. Step #3 - Sulfide (anion) means S2¯. Step #4 - because the formula must have a zero total charge, you write the formula Fe2S3. Example #4 – type the formula: tin(IV) phosphide The first character is Sn from the name tin. Roman numeral IV gives +4 as a tin charge. Phosphide provides P3¯. The formula for this compound is Sn3P4. This graph summarises the example #4: Practice Problems Write the correct formula: Answers 1) iron (II) chloride 2) copper (I) sulfide 3) lead (IV) iodide 4) tin (II) fluoride 5) mercury (I) bromide 6) nickel II) oxide 7) chromium (III) oxide 8) gold(I) iodide 9) manganese (II) nitride 10) cobalt (III) phosphide phosphide11) iron (III) chloride 12 copper (II) with 13lfide) lead(II) bromide 14) tin(IV) iodide 15) mercury(II) fluoride 16) nickel (IV) 15 oxide 17) manganese (III) chloride 18) chromium (II) nitride 1 9) gold (III) oxide 20) cobalt (II) phosphide 21) tin(II) (22) Mercury(I) sulphide 23) gold (III) bromide 24) manganese (II) oxide 25) 25) chloride 26) lead(IV) nitride 27) cobalt (III) oxide 28) copper(II) iodide 29) tin (IV) fluoride 30) iron(II) phosphide 1) FeCl2 2) Cu2S 2 3) PbI4 4) SnF2 5) Hg2Br2 6) NiO 7) Cr2O3 8) AuI 9) Mn3N2 10) CoP 11) FeCl3 12) CuS 13) PbBr2 14) S nio2 17) MnCl3 18) Cr3N2 19) Au2O3 20) Co3P2 21) SnS 22) Hg2S 23) AuBr3 24) MnO (25) CrCl2 26) Pb3N4 27) Co2O3 28) CuI2 29) SnF4 30) Fe3P2 binary compounds with variable tax common naming system Binary compound is made of two different elements. There may be one of each item, such as CuCl or FeO. There may also be several items for each item, such as Fe2O3 or CuBr2. This tutorial shows you how to name binary compounds (using a common naming system) from a formula when a variable fee vocation is involved. All four above formulas are examples of this type. It is important to remember: the lessons participating in this lesson have variable fees. The anions involved have only one charge. Antoine Laurent Lavoisier (1743-1794) reformed chemistry in the late 1700s, releasing the Méthode de nomenclature chimique (together with three co-authors) and Traité élémentaire de Chimie in 1789. He is known as the father of Modern Chemistry. Two typical chemical names up to this point in history are cut out stone land and phlogisticated vitriolic acid. There were hundreds of such names. One of Méthode's goals was to create chemical names based on chemical composition. Lavoisier's solution, which will be studied in this lesson, was to use different clasp to indicate the differences in composition. In particular, the use of -ous and -ic will be investigated. Here's what IUPAC currently says about this naming system: These systems are used, but not recommended: The system for naming suffixes -ous and -ic attached to the name of the calling name can be saved for elements with no more than two configurations. Example #1: FeO #1 - The first part of the name is the root and suffix of the first formula element. The use of iron root is ferr-. The clasp will be -ous or -ic. Here's how to set up a clasp. 1 Multiply the anion (O) charge by its subscript. Ignore the fact that it is negative. 2. Divide the result by the lower index (Fe). This gives a positive fee for calling. 3) the lower of the two values of a given calling is assigned to the ending-ous, and the higher the end -ic. The result of (1) and (2) a little higher is two. (As you memorize the various fees, you also internalize the above three steps.) This last part deserves a repeat: the bottom of the two values will use -ous end and higher will use -ic. I see you saying to yourself: As in the world I know which one is which one is higher? The answer is that you will know from your studies which one is smaller, which is higher. For example, iron takes on the +2 value and the +3 value. When you start learning these values, the lower and higher question becomes much easier. Trust me! Step #2 - anion named after the usual stem way plus ide. The answer to this example is iron oxide. Example #2: Fe2O3 When you multiply anion charge (negative two) from its subscript (three) and drop the sign, you get six for the answer. Then divide six of the two (iron sub-index) and you will get three. This means that the charge of each iron is positive three. Since this is the higher of the two charges, the term ferric is used. The answer to this example is iron oxide. Example #3: CuCl2 The first part of the name is from the roots of the first element: cupr-. Two chlorides are equal to -2, so Cu must be +2. -ic is used because +2 is higher of the two taxes copper is known to. The second part of the name is obtained from the root plus ide of the second character, so chlorine + chloride = chloride. This compound is named cupric chloride. Example #4: SnO The first part of the name is from the roots of the first element: stann-. One oxygen = negative 2, so one tin equals +2. The tin should be equal to +2, because you need to create a formula with zero total payments. -ous is used because +2 is the lower of the two tax tins are known to. The second element is oxygen (from the symbol O), so the name is taurus + ide = oxide. This compound is named stannous oxide. Practice Problems Write the correct name for: Answers: 1) NiS 2) PbBr4 3) Pb3N2 4) Fe2S3 5) FeI2 6) Hg2F2 7) Cu2S 8) SnCl2 9) HgO 10) Sn3P4 11) NiS2 12) PbCl2 13) Sn3N4 14) FeS 15) FeBr3 16) HgF2 17) CuS 18) SnI4 19) Hg2O 20) Pb3P4 21) NiO 22) SnO2 23) PbO2 24) Fe2O3 25) CuI 26) Hg2Cl2 27) CuO 28) Sn3N2 29) Hg2O2 30) CuCl2 1) nickelous sulfide 2) plumbic bromide 3) plumbous nitride 4) ferric sulfide 5) ferrous iodide 6) mercurous fluoride 7) cuprous sulfide 8) stannous chloride 9) mercuric oxide 10) stannic phosphide 11) nickelic sulfide 12) plumbous chloride 13) stannic nitride 14) ferrous sulfide 15) ferric bromide 16) mercuric fluoride 17) cupric sulfide 18) stannic iodide 19) mercurous oxide 20) plumbic phosphide 21) nickelous oxide 22) stannic oxide 23) plumbic oxide oxide 25) cuprous iodide 26) mercurous chloride 27) cupric oxide 28) stannous nitride 29) mercurous peroxide 30) cupric chloride Binary Compounds of Cations with variable fees Given name, Write formula common name system example #1: cuprous chloride #1 step - cuprous is a very specific calling name. It's Cu + and nothing more. I see you saying to yourself: How in the world do I know that cuprous means cu+? The answer - in your studies you know which vocation goes with Name. If you give you enough time for your studies, you will remember all the names and taxes as you have to memorize many other things in your life. Step #2 - chloride is the name of a specific anion. This is cl¯. Step #3 - remembering that the total formula tax should be zero, you write the formula CuCl. For example#2: Iron oxide Iron means Fe2+. Oxide means O2¯. Follow the usual rules, you write a FeO formula. Example #3: Iron sulphide Ferric gives Fe3 +. Sulfide is S2¯. The formula is a Fe2S3 sample #4: stannic phosphide Stannic means Sn4+. As an apple or bulb means specific things Phosphide means P3¯. Sn3P4. #5: Mercury chloride Mercury stands for Hg22+ Chloride means Cl. Hg2Cl2. This formula does not decrease. Practice Problems Write the correct name for: Answer 1) NiS 2) PbBr4 3) Pb3N2 4) Fe2S3 5) FeI2 6) Hg2F2 7) Cu2S 8) SnCl2 9) HgO 10) Sn3P4 11) NiS2 12) PbCl2 13) Sn3N4 14) FeS 15) FeBr3 16) HgF2 17) CuS 18) SnI4 19) Hg2O 20) Pb3P4 21) NiO 22) SnO2 23) PbO2 24) Fe2O3 25) CuI 26) Hg2Cl2 27) CuO 28) Sn3N2 29) Hg2O2 30) CuCl2 1) nickelous sulfide 2) plumbic bromide 3) plumbous nitride 4) ferric sulfide 5) ferrous iodide 6) mercurous fluoride 7) cuprous sulfide 8) stannous chloride 9) mercuric oxide 10) stannic phosphide 11) nickelic sulfide 12) plumbous chloride 13) stannic nitride 14) ferrous sulfide 15) ferric bromide 16) mercuric fluoride 17) cupric sulfide 18) stannic iodide 19) mercurous oxide 20) plumbic phosphide 21) nickelous oxide 22) stannic oxide 23) plumbic oxide 24) ferric oxide 25) cuprous iodide 26) mercurous chloride 27) cupric oxide 28) stannous nitride 29) mercurous peroxide 30) cupric chloride Binary Compounds The binary compound of the two nonmetals Greek system is one of two different elements. There may be one of each element, such as CO or NO. There may also be several items for each item, such as BF<SUB3 or OCl2. This tutorial shows how to name binary compounds from a formula when two non-metals are involved. All four above formulas are examples of this type. It is important to remember: there are no metals (which act as a vocation) involved. This means that one of the non-metals will be running a positive role, while the other is negative. Actually, you don't even need to know the fees, because the formula is directly from the cell names and their prefixes. Note that this lesson uses the heavy use of Greek number prefixes. Here is the first ten: one mono- six- two- seven hepta- three tri-eight octa- four tetra-nine nona- five penta- deca- #1 - write the name of N2O. For example#2 - Type a name for NO2. Step #1 - Part of the name is the unchanged name of the first formula element. In the examples above, this would be nitrogen. If the subscript for the first item is 2 or more, you add a prefix to the name. In the first example above, write dinitrogen. If the subscript is as shown in the second example above, you do not use the prefix. You just write a name, in this example it would be nitrogen. Step #2 - anion named after the usual stem way plus ide. In addition, a prefix is attached. In the first example, the prefix is mono-, because there is one oxygen. In the second example, use di- for two oxygen. The correct names of the two samples are dinitrogen monoxide and nitrogen dioxide. Note that monoxide is written, not monooxide. It sounds better when it comes out loud. Example #3 - Type a name for IF7. Step #1 - the first element is iodine and there is only one. This part of the name will be iodine, NE monoidine. Step #2 - the second element is fluorine, so fluoride is used. Since there are seven, the prefix hepta is used. The name of this compound is iodine heptafluoride. Example #4 - Type the N2O5 name. Step #1 - the first element is nitrogen and there are two. This part of the name will be dinitrogen. Step #2 - the second element is oxygen, so oxide is used. Since there are five, the prefix penta is used. The name of this compound is dinitrogen pentaxide. Many write as dinitrogen pentoxide. ChemTeam believes that both are considered correct, but the second must be preferred. Example #5 - Type a name for XeF2. The first part of the title is from the name of the first item: xenon. Since there is only one atom, the prefix is not used. The second part of the name is from the root of the second character plus ide, as well as the prefix di-, so di + fluorine + ide = difluoride. This compound is named xenon difluoride. Example #6 - Type the N2O4 name. The first part of the name is from the name of the first element: nitrogen. Since there are two atoms, the prefix di- is used to provide dinitrogen. The second part of the name is from the root of the second character plus ide, as well as the prefix tetra-, so tetr + ox + ide = tetroxide. This compound is named dinitrogen tetroxide. Pay attention to a flushing tetra. Just a reminder: this naming system certainly does not have a substantially acceptable name, but is often referred to as the Greek system (or method). This includes the use of Greek prefixes to name binary compounds of two non-metals. Sometimes you will see a stock system for these types of compounds. Here's what IUPAC is currently saying about this practice: Stock notation can be applied to both callers and anions, but that it should not apply to compounds between non-metals. Practice Problems Write the correct name for: Answers: 1) As4O10 2) BrO3 3) BN 4) N2O3 5) NI3 6) SF6 7) XeF4 8) PCl3 9) CO 10) PCl5 11) P2O5 12) S2Cl2 13) ICl2 14) SO2 15) P4O10 16) UF6 17) OF2 18) ClO2 19) SiO2 20) BF3 21) N2S5 22) CO2 23) SO3 24) XeF6 25) KrF2 26) BrCl5 27) SCl4 28) PF3 29) XeO3 30) OsO4 1) tetrarsenic decoxide 2) bromine trioxide 3) boron nitride 4) dinitrogen trioxide 5) nitrogen triiodide 6) sulfur hexafluoride 7) xenon tetrafluoride 8) phosphorous trichloride 9) carbon monoxide 10) phosphorous pentachloride 11) diphosphorous pentoxide 12) disulfur dichloride 13) iodine dichloride 14) sulfur dioxide 15) tetraphosphorous decoxide 16) uranium hexafluoride 17) oxygen difluoride 18) chlorine dioxide 19) silicon dioxide 20) boron trifluoride 21) dinitrogen pentasulfide 22) carbon dioxide 23) sulfur trioxide 24) xenon hexafluoride 25) krypton difluoride 26) bromine pentachloride 27) sulfur tetrachloride 28) phosphorous trifluoride 29) xenon trioxide 30) osmium tetroxide Compounds Involving Polyatomicsion These compounds to be observed are not binary compounds. They contain three or more elements, not just two binary compounds. The Greek method will not be used. This naming technique is used only for two non-metal binary compounds. This means that if you see a formula like BaSO4, the name is not barium monosulfure tetraoxide. Many ignorant ChemTeam students over the years have made this mistake and have suffered for it. Therefore, warning: it is important that you learn to recognize the presence of polyatomic in the formula. Many ChemTeam students made their first priority to make a set of flashcards with a name on one side and his charge on the other. Then carry them everywhere and use them. The fees used will be a combination of fixed charges and variable charges. You need to know which ones. Another caveat: you also need to know the fees associated with each polyatomicoon. For example, NO3 is called nitrate and has minus one charge. Once again, many ignorant ChemTeam students thought this meant nitrates have a minus three fee. IT'S NOT. Using parentheses When more than one polyaminal is required, the brackets are used to add a washbas in brackets with a lower index when going outside the bracket. For example, the first formula used is Fe(NO3)2. This means that two NO3 compounds are involved. In bracket formula would be FeNO32, far from the correct formula. When you say a word aloud that includes a parenthesis, you use the word taken, as in the ammonium sulfide formula that is (NH4)2S. Loudly, you say: N H four taken twice S. OR copper (II) chlorate formula, which is Cu(ClO3)2. You say: Cu Cl O three taken twice. Example #1 - write the name of the Fe(NO3)2 action #1 - decide if the calling is displayed Fee. If so, you'll need Roman numerals. In this scenario, the iron indicates a variable charge. If a variable fee vocation is involved, you must set a Roman number. You can do this by calculating the total fee deposited by polyatomics. In this case, NO3 has minus one charge and there are two of them, bringing the total to minus 2. Therefore, the iron must be a positive two so that the total formula tax is zero. Step #2 - set the name of the polyaminal. Nitrate is NO3¯. The correct name is iron (II) nitrate. The common name would be iron nitrate. Example #2 - type the name of the NaOH step #1 - the vocation, Well+, does not show a variable fee, so you do not need Roman numerals. The name is sodium. Step #2 - OH¯is recognized as hydroxide. The name of this compound is sodium hydroxide. Usually, there is a cry at this time in chemteam class. But how do you know that OH¯is hydroxide? is a simple wail. Stock ChemTeam's answer is well, how do you know someone? How do you know your phone number? How do you know the name of your best friend? Actually, how do you know your name? You must memorize three things: name (hydroxide), symbol (OH) and boot (minus one). You have to put time to learn this nomenclature stuff. It's not easy, and ChemTeam understands that you'd rather spend your time doing more important things: going to great places with friends, spending time with members of the opposite sex, spending your parents' money, sleeping, and so on. Maybe next time. Now let's move on. Example #3 - type KMnO4 #1 action name - vocation, K+, does not display a variable fee, so you don't need Roman numerals. The name is potassium. Step #2 - MnO4¯, which is recognized as a permanganato ion. The name of this compound is potassium permanganate. Solution #4 – Type the cu2SO4 step name #1 decide whether the calling is a variable fee. If so, you'll need Roman numerals. In this scenario, the copper displays a variable charge. If a variable fee vocation is involved, you must set a Roman number. You can do this by calculating the total fee deposited by polyatomics. In this case, SO42 has minus two charges and there is only one, so a total of minus 2. Copper must therefore be positive. Why? Well, there must be a positive two to go with the negative two in order to make zero. Since the formula shows two copper atoms involved, each must be plus one charge. Step #2 - set the name of the polyaminal. Sulfate is SO42¯. The correct name is copper (I) sulphate. The common name would be cuprous sulfate. Practice problems Write the correct name: Answers: 1) AlPO4 2) KNO2 3) NaHCO3 4) CaCO3 5) Mg(OH)2 6) Na2CrO4 7) Ba(CN)2 8) K2SO4 9) NaH2PO4 10) NH4NO3 11) Sn(NO3)2 12) FePO4 Cu2SO4 14) Ni(C2H3O2)2 15) HgCO3 16) Pb(OH)4 17) Cu2Cr2O7 18) Cu(ClO3)2 19) FeSO4 20) Hg2(ClO4)2 21) KClO3 22) SnSO4 23) Al(MnO4)3 24) Pb(NO3)2 25) Mg3(PO4)2 26) cuH2PO4 27) CaHPO4 28) Fe(HCO3)3 29) Na2CO3 30) MnSO4 1) aluminium phosphate2) potassium nitrite 3) sodium hydrogen carbonate [sodium bicarbonate] 4) calcium carbonate 5) magnesium hydroxide 6) sodium cromate 7) barium cyanide 8) potassium sulphate 9) sodium salt dianodenil phosphate 10) ammonium nitrate 11) tin (II) nitrate [stannous nitrate] 12) iron(III) phosphate [iron phosphate] 13) copper (II) sulphate [cup sulphate] 14) nickel (II) acetate [nickel acetate] 15) mercury(II) carbonate [mercury carbonate] 16) lead (IV) hydroxide [plumb hydroxide] 17) copper (I) dichromate [copper dichromate] 18) copper (II) chlorate [copper chlorate] 19) iron (II) sulphate [iron sulphate] 20) mercury (I) perchlorate [mercury perchlorate] 21) potassium chlorate 22) tin (II) sulphate [stannous sulphate] 23) aluminium permanganate 24) lead(II) nitrate [sanitary nitrate] 25) magnesium phosphate 26) copper(I) divane phosphate 27) calcium hydrogen phosphate 28) iron(III) hydrogen carbonate 29) sodium carbonate 30) manganese (II) sulphate [mangan sulphate] Nomenclature of inorganic acids Recognising acid at this beginning, you will recognize acid according to the fact that its formula starts with H, as in the following examples: HCl HNO3 H2SO4 HClO3 H3BO3 How you become more complex in your chemistry You will realize that there are many acid formulations that do not start with H, but they will almost all be left for another time. There is one exception: the formula CH3COOH must be recognized as acetic acid. A special way, as it is written, is common in organic chemistry. An alternative way to write acetic acid is HC2H3O2. This is done in an inorganic style, which you are currently studying. Last comment before viewing how to name acid: the formula H2O should not be considered acid. It's a water formula. It's not acid. When you study acid and base behavior later in the school year, you will learn more about the role of water in acid and base chemistry, but not now. Naming acids The sequence of acid names, HCl, HClO, HClO2, HClO3 and HClO4 will be discussed in order. HCl is a binary acid. All binary acids are named in the same way: 1. the prefix hydro is used. 2. the root of the anion is used. 3) the res and shed is used. 4) The word acid is used as the second word in the title. The name of HCl is hydrochloric acid. Other binary acids for which you are responsible are HF, HBr, HI and H2S. 1) HClO is an acid containing polyatomics. In the formula, you must recognize the polyatomic. There is no other way to find out the name. If you do not recognize polyatomic, then you drowned without a trace. Polyatomics are and its name is hypochlorite. Every time you see it is enough, you can change it to ous and add the word acid. The name of HClO is hypochloric acid. 2) HClO2 has ClO2¯polemic. The name of shijon is chlorite. Because the apfix is used, it is changed to ous. The name of HClO2 is chloric acid. 3) HClO3 contains ClO3 polyatomic ion and its name is chlorate. Any time you know that you ate a limb used polyatomic, you use ic when its acid. The name HClO3 is chloric acid. 4) HClO4 has ClO4 polyatomic john and its name is perchlorate. Since used to eat the apag, it is replaced by ic. The name HClO4 is perchloric acid. According to ChemTeam, there are two keys. You must: 1. recognize when polyatomics are located, and 2. know its name. Only then can you know that change it suffix ous and ate suffix ic when it is acid. Practice issues 1) H3PO4 2) H2CO3 3) H2SO4 4) HIO3 5) HF 6) HNO2 Write the formula for the following acids: 7) hydro bromide acid 8) hydrocyanic acid 9) nitric acid 10) sulphuric acid 11) phosphoric acid 12) acetic acid 1) H3PO 3 2) HClO3 3) H2SO3 4) HBrO3 5) HI 6) CH3COOH Write the formula for the following acids: 7) hydrochloric acid 8) hydrochloric acid [this must turn] 9) perchloric acid 10) hydrochloride 1) phosphoric acid 2) carbonic acid 3) sulphuric acid 4) iodine acid 5) hydrofluoric acid 6) nitric acid 7) HBr 8) HCN 9) HNO3 10) H2SO3 11) H3PO3 12) CH3COOH or HC2H3O2 1) phosphorus acid 2) chloric acid 3) sulphuric acid 4) bromocid 5) hydrochloric acid (not hydrocytic acid!) 6) acetic acid 7) HCl 8) H2S 9) HClO4 10) HOH (commonly known as H2O or water) 1. ammonium sulphide 26. nickel (II) iodide 2. sodium nitrate 27. mercury oxide 3. military bromide 28. lead (II) chlorite 4. aluminium sulphate 29. hydrogen iodide 5. potassium nitrate 30. iron (II) bisulphite 6. iron carbonate 31. magnesium nitrate 7. lead (II) phosphate 32. iron (III) chromate 8. diphosphorus pentoxide 33. iron (II) chromate 9. coffee hydroxide 34. copper (II) hydroxide 10. calcium fluoride 35. cuprous carbonate 11. nickel (II) nitrate 36. chromium acetate 12. silver cyanide 37. calcium chlorate 13. ammonium sulphite 38. ammonium oxide 14. zinc sulphate 39. aluminium perchlorate 15. tin (II) chloride

40. zinc bicarbonate 16. antimononium (III) chloride 41. sodium phosphate 17. silver sulphide 42. silver hypochlorite 18. magnesium hydroxide 43. ammonium phosphate 19. Ammonium 44. iron chlorite 20. nickel (II) acetate 45. potassium sulphide 21. sodium chromate 46. tin (IV) bromide 22. chromium bisulphate 47. lithium chromate 23. potassium permanganate 48. magnesium bisulphate 24. silver perchlorate 49. iron phosphate 25. potassium phosphate 50. calcium sulphate dihydrate 51. aluminium acetate 76. lithium hypochlorite 52. calcium chloride dihydrate 77. oxygen difluoride 53. barium chromate 78. cobalt (II) hydrogen sulphate 54. cobalt chloride 79. acetic acid (see #128) 55. barium chloride dihydrate 80. barium hypochlorite 56. sulphuric acid 81. ammonium hydroxide 57. potassium hydroxide 82. cobalt (II) iodide 58. zinc bisulphite 83. chromium (II) bicarbonate 59. sodium sulphite 84. sodium hydroxide 60. cobalt sulphate 85. silver nitrate 61. iron oxide 86. mercury (II) nitrate 62. silver phosphate 87. hydrochloric acid 63. sodium hypochlorite 88. aluminium bisulfite 64. ammonium chromate 89. cobalt (III) hydrogen sulphate 65. barium carbonate 90. iron hydrogen carbonate 66. calcium iodide 91. phosphorus pentabroide 67. copper sulphate 92. nickel (II) chloride hexahydrate 68. cuprous chloride 93. ammonium aluminium sulphate 69. iron carbonate 94. iron (III) hydrogen carbonate 70. zinc phosphate 95. mercury(I) hydrogen phosphate 71. sodium nitrite 96. plumbin hydrogen carbonate 72. silver oxide 97. mercury hydrogen carbonate 73. nickel (II) bromide 98. mercury hydrogen phosphate 74. magnesium oxide 99. copper (II) sulphate pentahydrate 75. mercury perchlorate 100. chromium divadrogen phosphate 1. (NH4)2S 26. NiI2 2. NaNO3 27. Hg2O 3. CuBr2 28. Pb(ClO2)2 4. Al2(SO4)3 29. HI 5. KNO3 30. Fe(HSO3)2 6. FeCO3 31. Mg(NO3)2 7. 32. Fe2(CrO4)3 8. P2O5 33. FeCro4 9. Cu(OH)2 34. Cu(OH)2 10. CaF2 35. Cu2CO3 11. Ni(NO3)2 36. Cr(C2H3O2)3 12. Agcn 37. Ca(ClO3)2 13. (NH4)2SO3 38. (NH4)2O 14. ZnSO4 39. Al(ClO4)3 15. SnCl2 40. Zn(HCO3)2 16. SbCl3 41. Na3PO4 17. Ag2S 42. AgClo 18. Mg(OH)2 43. (NH4)3PO4 19. (NH4)2CO3 44. Fe(ClO2)2 20. Ni(C2H3O2)2 45. K2S 21. Na2Kro4 46. SnBr4 22. Cr(HSO4)3 47. Li2Cro4 23. KMnO4 48. mg(HSO4)2 24. AgClo4 49. Fe3(PO4)2 25. K3PO4 50. CasO4.2H2O 51. Al(C2H3O2)3 76. LiClo 52. CaCl2.2H2O 77. Ip 2 53. BaCro4 78. Co(HSO4)2 54. CoCl3 79. CH3COOH 55. BaCl2.2H2O 80. Ba(Clo)2 56. H2SO3 81. NH4OH 57. KOH 82. CoI2 58. Zn(HSO3)2 83. Cr(HCO3)2 59. Na2SO3 84. Naoh 60. Coso4 85. AgNO3 61. Fe2O3 86. Hg(NO3)2 62. Ag3PO4 87. HCl 63. NaClo 88. Al(HSO3)3 64. (NH4)2Kro4 89. Co(HSO4)3 65. BaCO3 90. Fe(HCO3)3 66. CaI2 91. PBr5 67. 92. NiCl2.6H2O 68. CuCl 93. NH4Al(SO4)2 69. Fe2(CO3)3 94. Fe(HCO3)3 70. Zn3(PO4)2 95. Hg2HPO4 71. NaNO2 96. Pb(HCO3)4 72. Ag2O 97. Hg(HCO3)2 73. NiBr2 98. Hg2HPO4 74. MgO 99. CusO4.5H2O 75. Hg(ClO4)2 100. Cr(H2PO4)3 Cr(H2PO4)3

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