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## Equilibrium worksheet answers pogil

Redirect to download Balance Pogil Answer Key PDF after seconds Balance 1 Balance At what point is a reversible reaction completed? Why? Most of the reactions we've studied this year have been forward reactions—once the reactant has changed to the product that stays that way. We can assume that eventually the reaction will end when the restrictive reaction is over. However, in a reversible reaction (the one that can take place in both directions), once the product is formed it can be converted back to the reactive and a continuous cycle occurs. When do reactions like these end? Is there ever an end, and can we predict how reactive or product will be present when the reaction is done? Model 1 – Reversible reaction  $A(g) \rightleftharpoons B(g)$  1. What's reacting to the reaction to model 1? 2. What is the product in reaction to model 1? 3. What is the significance of the double arrow in the equation in model 1? 4. Imagine that the reaction to Model 1 starts with 100 A molecules and zero B molecules. Explain why the concentration of substance A will never reach zero. 5. Imagine that the reaction to Model 1 starts with zero A molecules and 100 B molecules. Explain. 2 POGIL™ Activities for High School Chemistry 6. Consider an initial concentration of 5.00 moles of A and zero moles of B for reaction to model 1. If 60% of the available A molecules react every minute, calculate the concentration of A and B after one minute. Fill in the table below with your answers. Be prepared to discuss your calculation method with the class. A Initial Starting Moles of A 5.00 Moles 0.00 Home Start Moles of B Change Forward Moles of A Reacted - • → + Change Forward Moles of B did End Moles of A after 1 minute = End Moles of B after 1 minute 7. Did you use the mole ratio between substance A and substance B in question 6? If not, make a correction. 8. Consider an initial concentration of 5.00 moles of A and 2.00 moles of B for reaction to model 1. If 60% of the available A molecules react every minute and 20% of the available B molecules also react every minute, calculate the concentrations of A and B after one minute. Fill in the table below with your answers. Hint: Since forward reactions and reversal occur simultaneously, calculate the reverse change based on the original 2.00 moles of B. A Initial Starting moles of A 5.00 moles 2.00 Initial Start moles of B Change Forward Moles of A reacted - • → + Forward Moles of B did Change Reverse Moles of A reacted + ← - Change Reverse Moles of B reacted End Moles of A after 1 minute = End Moles of B after 1 minute 9. Get a set set conditions by the instructor. Enter the original moles of A and B in model 2, as well as the percentage of A molecules that react every minute (above the first forward arrow) and the percentage of B molecules that react every minute (above the first inverse arrow). Balance 3 Model 2 – Achieve Balance A B 1 minute Initial Change Forward - • → + Reverse Change + ← - End = = 2 minutes Change forward - • → + Change Reverse + ← - End = 3 minutes Change Forward - • → + Change Reverse + ← - End = 4 minutes Change Forward - • → + Change Reverse + ← - End = 5 minutes Change Forward - • → + Reverse Change + ← - End = 6 minutes Change forward - • → + Change Reverse + ← - End = = 7 minutes Change forward - • → + Change Reverse + ← - End = = 8 minutes Change forward - • → + Change Reverse + ← - End = = 9 minutes Change forward - • → + Change Reverse + ← - End = =

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