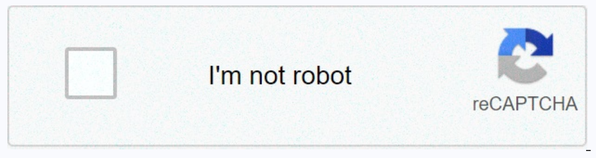

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Chemical bond lab answers

Redirect to download PDF chemical link lab responses after seconds Thank you for your participation! Introduction The purpose of the first part of the experiment is to be able to understand binding by observing the physical properties of molecules. The purpose of the second part of the experiment is to be able to use quantum mechanics methods in a program called Web MO to draw the Lewis structure of the different compounds in the lab. Followed by the calculation of the structure and properties of different molecules. A chemical compound is made up of atoms that are united by the attractive forces of chemical bonds. These bonds can be ionic, covalent, or metallic. A ionic bonus it's a kind of bonus that generates two ions, in which the distribution of electrons is transferred. The ion union implies the union between metals and non-metallic metals, and an example of the type of compound involved in this type of the binding would be potassium chloride (KCl). A covalent bond it's a kind of bond in which there's a balance of repulsive and attractive forces between atoms as they share electrons. Covalent union involves the union between non-metallic and/or metalloids, and an example of the type of element involved in this type of bond would be hydrogen (H 2). A metal joint it's a link in that many electrons separated between many positive ions are grouped together. Metal joint involves the union between metals, and an example of the type of element involved in this type of union would be copper (Cu). The type of bonding in a compound can be determined by the physical properties of a substance, such as its conductivity, solubility in water, and melting point. Valence Shell's electron pair repulsion theory, A.K.A VSEPR, is a model in chemistry used to predict the geometry of individual molecules from the given amount of pairs of electrons surrounding the atom. VSEPR will help us associate the electron configuration with molecular geometry. Hydrides involve a negatively charged hydrogen atom bonding to a more positive element or group. Isoelectronic molecules involve two atoms containing the same number of valence electrons and have the same electronic structure. In the second part of the hydrides and isoelectronic molecules, lewis structures of hydrides and isoelectronic molecules will be created laboratory using the WebMO program. In Part I of the laboratory, the main focus is experiments with respect to the melting point of a compound and solubility in distilled water. From that data will be collected to discover the information needed to determine whether compounds are ionic or molecular. In Part II of the laboratory, energy minimized three