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Chemical reactions test review answers

Chemical Reaction Test Review Answers | Updated 6558 kbps 4180 Chemical Reaction Test Review Answers | Full 4323 kbps 6103 Chemical Reaction Test Review Answers [Most Popular] 2586 kbps 40624 Chemical Reaction Test Review Answers | New 5308 kbps 36416 Chemical Reaction Test Review Answers | added by users 5984 kbps 18212 Problems with the test review. The teacher did not give us answers. It would be easier for me to learn with answers. 1.) What is the percentage (in weight) of hydrogen in acetic acid, HC₂H₃O₂?.) How many H atoms are there in 4.9 g of water? 3.) What is the percentage (by weight) of nitrogen in Ca(NO₃)₂? What is an empirical formula for a compound containing 85.63% C and 14.37% H? 5.) how many moles H₂O produced when consuming 0.30 moles O₂? 6.) how many moles H₂O produced when consumed 4.3 gram O₂? 7.) What is the percentage (in weight) of oxygen in C₈H₈O₃? 8.) Calculate formula weight (FW) ammonium nitrates. 9.) Calculate the mass in grams 0.333 moles CO₂. 10.) Convert 100.0 grams of barium hydroxide into a mole sample. 11.) How many chloride ions are present in 50.0 g of iron chloride (ii) chloride? 12.) Calculate the mass in the grams of the sample containing water molecules 1.85 x 10³⁴. 13.) Which of the following actions may not be an empirical formula? A) C₂H₂ B) CH₂ C) CH₃ D) C₂H₃O₄. The empirical acid formula is CH₂O and the molar mass is 360. What is a molecular formula? 15.) When 0.125 moles of butane, C₄H₁₀, are burned by excess oxygen, ideally how many CO₂ moles are formed?

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