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Combustion definicion corrosion

Examples of reactions of interestIt is a lot of chemical reactions of interest, either because of their usefulness, scope, influence, danger, etc. In this section we will discuss the following: Photosynthesis. Digestion and breathing. Acid rain. Metal corrosion. Combustion. Industrial getting of substances. Decomposition of ozone. Electrochemical reactions. Petrochemical industry. The scene on the right speaks to some aspects of the above reactions. Page 2Th in this two weeks you will learn how: Start with chemical stachytherapy. Correctly understand and use the basic sizes and laws necessary to perform calculations on chemical reactions. Get to know and understand the relationships between chemical sizes associated with sesichiometry. Granddad information from a certain chemical reaction. Perform calculations of mass and volume from chemical reactions. Understand the concept of chemical reaction performance. He knows different ways to classify chemical reactions. Know some chemical reactions of interest. Assess the importance of chemical reactions, and therefore the need for their study and knowledge. Page 3 Activities to send to the tutor These activities can be printed from here Combustion is a chemical process of rapid oxidation that is accompanied by a low energy shift in the form of heat and light. In order for this process to take place, it is necessary to have fuel, fuel and heat. A material capable of burning and combining with oxygen is known as fuel. In ordinary combustion, fuel is a composite substance, such as hydrocarbons (petroleum gas, gasoline, kerosene, paraffin, etc.). Oxygen, an essential element for the oxidation process that occurs and continues, is known as oxidant. What is CombustionSe defines as rapid oxidation chemical reaction combustion, to one that is accompanied by a low release of energy in the form of heat and light. In order for this process to take place, it is necessary to have fuel, oxidant and heat. Any substance that can burn under certain conditions is called combustion. Like any thing that can burn or suffer rapid oxidation. Types of combustion Reaction created by combustion components; Fuel material and oxidants, is detected in three types of reactions, which are as follows:Full combustionThis combustion reacts when the combustible material is completely oxidized and consumed, and then other oxygenated compounds are insufficiently overdosed, such as sulfur dioxide, carbon dioxide or water vapor. Squiometric combustIt is the name given to complete combustion, they occur when methane is converted into CO2 and H2O, use the right amounts of oxygen for their reaction and usually occur only in a controlled environment of the laboratory, using the necessary instruments, for example in the case of methane powder used a tablespoon of flue. Incomplete combustion Of those in which half of the oxidized compounds (also called unextended) appear from flue gases, such as carbon monoxide (CO), hydrogen, carbon particles, etc. Combustion process Fuel must reach a minimum temperature to burn, this temperature is the so-called ignition point or the temperature of inflammation. Flammable materials have a low temperature of inflammation and easily get into combustion. If coal or sulfur is burned in equal quantities, it will be noticed that the caloric energy from coal is greater than the energy of sulfur. This means that fuels, when burned, do not release equal amounts of heat. There are some that give a lot of calories of energy, while others less heat. As a result of the process, products are obtained by combustion. This depends on the nature of the fuel, but water vapor, carbon dioxide and carbon are usually produced. The fact that burning fuel releases significant amounts of energy gives these materials special importance, as they can be used for our usefulness. Industries, factories and electricity generation plants use combustion to get the energy they need to function. Today, hydrocarbons are at the foreground among energy sources. Combustion productsHumols formed by solid and liquid particles suspended in the air. With sizes between 0.005 and 0.01 microns. It has irritating effects on the mucous membranes. Smoke is practically the first risk factor in the development of fire, before you feel the effect of increasing temperature. They are then:White smoke: burning plant products, animal feed, food, etc. Yellow smoke: chemicals containing sulfur, fuels containing hydrochloric acids and nitric acids. Gray smoke: cellulose compounds, artificial fibers, etc. Black smoke lights up: rubber. Dark black smoke: oil, acrylic fibers, etc. Smoke will also mix with toxic gases that will change its color: White smoke: it burns freely. Flame: The temperature varies depending on factors, such as fuel type and oxidant concentration. Heat: heat is a heavy form of energy, which will raise the temperature. Examples of combustionA wax candle: initially, the chemical reaction occurs only in the candlestick. however, after the flame reaches the wax, there is also a reaction in the wax. Burning wood: hydrocarbons in wood are combined with oxygen in the form of water and carbon dioxide. It's a very energetic reaction, so it generates large amounts of heat and light to release that energy. Flaming phosphorus: When phosphorus is rubbed against a slightly rough surface, friction generates such heat in the head of phosphorus (consisting of phosphorus and sulfur) to produce flames. This is an incomplete reaction because the remains of phosphorous wax paper remain. Burning coal: Burning coal reacts and turns from solid to gas. In this reaction, energy is released in the form of heat. Fireworks: When fireworks are lit, heat causes chemicals in it to react with oxygen in the atmosphere to produce heat and light. You could say it's an incomplete reaction. Bonfires: Campfires are examples of the type of reaction that occurs between dried leaves, paper, firewood or any other hydrocarbon and caloric energy load (such as a lit match or a spark created by stones). Gas stove: gas stoves work with propane and butane. These two gases, when in contact with the initial caloric energy load (for example phosphorus), are burned. This is a complete reaction, because it does not create waste, here it can cause spontaneous combustion. Wildfires: Wildfires are examples of uncontrolled reactions. As with firewood, these are incomplete reactions because they leave debris. Strong bases and organic matter: as for these materials such as causal soda, it reacts when it comes into contact with organic matter. Oil fires are spontaneous flames that form in swamps with a high content of decaying organic matter. Fuels in engines: the internal combustion engine is used in cars carrying hydrocarbons so that they can work inside the combustion chamber, to be gasoline one of the main components for this to be from the internal reaction. Combustion of methanol: also known as methyl alcohol, is an example of a perfect reaction, because it creates nothing but water and carbon dioxide. Burning metal magnesium: This is an example of a reaction in which water or carbon dioxide is not released. In this case, the product is magnesium oxide. It's incomplete combustion because it produces magnesium oxide. Explosives: Explosives, such as gunpowder and nitroglycerin, generate a combustion reaction and occur in milliseconds. It should be noted that there are weak and strong explosives. Gunpowder: Gunpowder is a weak explosive. In the case of low explosives, confined to confined spaces (such as a weapons camera) so that they can function. Images of combustion In addition, we will show you some images of combustion and different results obtained in each of them: FREQUENTLY asked questions about combustionOn is produced through a rapid oxidation chemical reaction that is accompanied by a low release of energy in the form of heat and light. In order for this process to take place, it is necessary to have fuel, oxidant and heat. Read more Widely used in devices that help move people from one place to another (cars, buses, planes, boats, etc.). Similarly, it is also used in homes to perform multiple functions, such as gas stoves or in kettles for cooking food, in candles sometimes used for lighting, etc. Read more They are the ones that consume fuel in the most violent way and, in addition to the high heat rate, create light. For example, a lit candle, match or fire. Read more To produce a chemical reaction you need fuel, oxidant and reach the so-called ignition temperature, that is, you need one burning element (fuel) and another that produces a reaction (oxidant) and usually oxygen in the form of gaseous O2. More readily Chemical reaction releases a large amount of energy in the form of heat (heat energy) which consequently leads to the expansion of gases (carbon gas and water vapour) creating flames, which is a mass of incandescent gas that reflects heat and light, and is in contact with a flammable substance. Read more

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