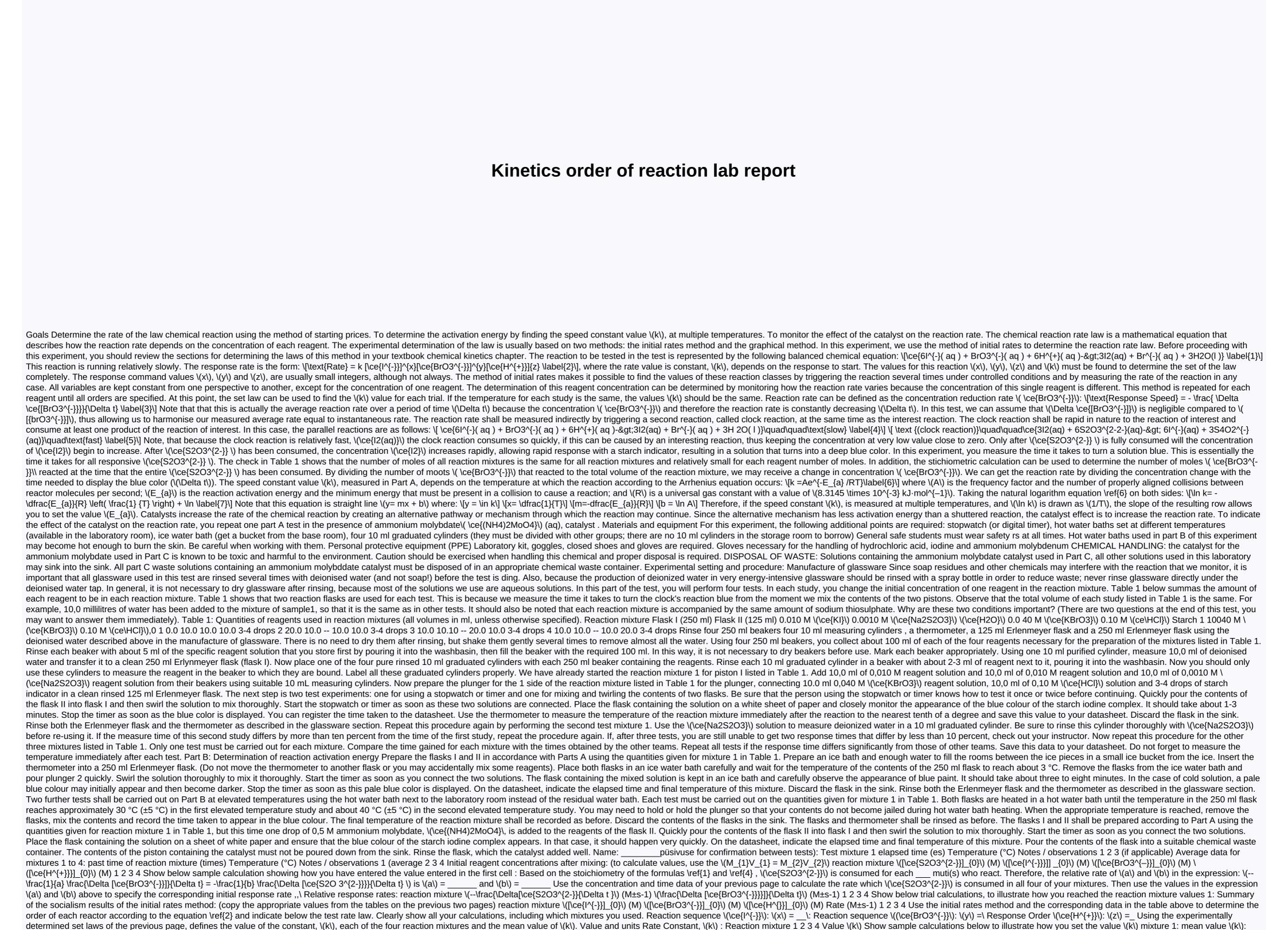
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Units \(k\): In this section we use the data collected in parts A and B of the mixture 1 to determine the activation energy value, \(E {a}\), for the reaction. Fill in the following table using the test data and calculated values obtained in part a and b of the test. Add your additional calculations separately on the back of

the laboratory report. Be sure to specify the speed constant value \(k\), for each temperature mentioned below, and give the appropriate units (the mean value specified in Part A \(k\) applies only to the room temperature record below). Summary of reaction activation energy determination results: Ice water room temperature *Approx. 30 °C About 40 °C Mixture temperature, \(T\) (°C) Mixture temperature, \(T\) (K) Elapsed time (id) \(--\frac{\Delta [\ce{BrO3^{2-}}]}{\toperature}\) (M±s-1) Constant speed and units, \(k\) \(\) \(\) (In k\) (no units) \(1/T\) (K-1) * Use part 1 to use the room temp. Your graph should have a suitable title and axes with the appropriate scope. Add the most appropriate line to the data you've drawn with the Excel trendline feature, and have Excel display the equation of that row and its R2 value. Send this graph with your report. Use this graph to determine the value of activation energy (\(E \ {a}\)\) and its reaction frequency factor \(A\) (be sure to add the correct units for each!). Let me see your calculations: \(E \ {a}\)\:

Questions The volume of reagents listed in Table 1 varies one by one, except for the volume of sodium thiosulphate, which is the same in each study. Why do we

keep the same amount of nathaniel in every study? In Part A, even if the concentrations of reactors are altered in each study, the experimentally determined values of the speed constant for each test should be quite similar. Why is that? How do the results in Part B also support your answer to the previous question? What is the qualitative relationship between the reaction rate and temperature obtained from the results in Part B? Can you give a physical explanation to explain this? What's going on at the molecular level? How does the presence of a catalyst affect the activation energy of this reaction based on the results of Part C? What additional test could you do to determine the value of the activation energy for the catalyzed reaction? What is one problem you could face if you performed the test you described in your answer to the previous question? How could you try to get over this problem? Let's say the student prepares a response mixture 2 (see Table 1 in Part A). mix the contents of reaction flasks 1 and 2, taking 75 seconds to appear blue. After mixing, the initial concentrations of the reactors are calculated. For example, the concentration of \(\ce{\sc{82O3^{2-}}}\) is calculated as follows: \(\[\ce{\sc{82O3^{2-}}}\] \(\ce{\sc{82O3^{2-}}}\] \(\ce{\sc{82O3^{2-}}}\) \(\ce{\sc{82O3^{2-}

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