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keep the same amount of nathaniel in every study? In Part A, even if the concentrations of reactors are altered in each study, the experimentally determined values of the speed constant for each test should be quite similar. Why is that? How do the results in Part B also support your answer to the previous question? What is the qualitative relationship between the reaction rate and temperature obtained from the results in Part B? Can you give a physical explanation to explain this? What's going on at the molecular level? How does the presence of a catalyst affect the activation energy of this reaction based on the results of Part C? What additional test could you do to determine the value of the activation energy for the catalyzed reaction? What is one problem you could face if you performed the test you described in your answer to the previous question? How could you try to get over this problem? Let's say the student prepares a response mixture 2 (see Table 1 in Part A). mix the contents of reaction flasks 1 and 2, taking 75 seconds to appear blue. After mixing, the initial concentrations of the reactors are calculated. For example, the concentration of S2O3^{2-} is calculated as follows: $\frac{[\text{S2O3}^{2-}]}{(0.010\text{ L}) + (0.0010\frac{\text{mol}}{\text{L}})((0.050\text{ L}))} = 2.0 \times 10^{-4}\text{ M}$ $\frac{[\text{BrO}_3^-]}{[\text{H}^+]} = \frac{[\text{BrO}_3^-]}{[\text{H}^+]}$ = calculate the initial response rate $(-\frac{\Delta [\text{BrO}_3^-]}{\Delta t})$ in $\frac{\text{molce}[\text{BrO}_3^-]}{\text{L}\cdot\text{dot s}}$. Let's say the time required to turn 1 blue is 145 seconds. Calculate the initial concentrations of the reagents after mixing reaction mixture 1 and use the initial rates method to determine the reaction sequence, $\ln(x)$, in relation to $\ln(\frac{[\text{H}^+]}{[\text{BrO}_3^-]})$.

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