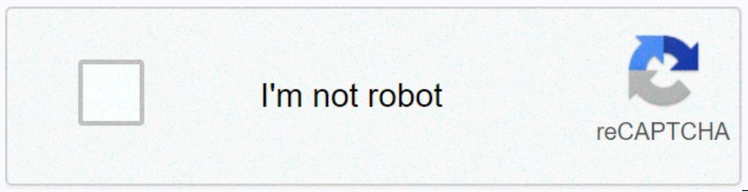

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Ib chem study guide

IB chemistry may not be as easy as this penguin makes it look. So to help you out, I compiled the best FREE online IB chemistry study manuals and notes into one useful article. I organized this IB Chemistry Study Guide using the order set out in Syllabics Chemistry IB. 2020 IB Chemistry Test Cancellation Due to COVID-19 Due to COVID-19 (coronavirus) pandemic, all IB exams for May 2020 have been cancelled and course dates have been extended for schools that have been closed. (Yeah – it definitely includes the IB Chem exam, too!) Stay up to date with the latest information on what this means for IB diplomas, course credit for IB classes and more with our ARTICLE FAQ IB COVID-19. How to use this IB Chemistry Study Guide If there is one specific topic you need more help with, use the Command+F feature on your computer to search for this guide for this topic. So, if you're hoping to read about mole concept, use Command+F to make a search function. Enter mole concept and it will bring all study materials for Mole Concept. If you're looking for aggregate material to help you study on IB chemistry papers, check out the overall reviews section for large overall study resources. I've given notes and tutorials by topic. You should check out this article during the school year to help you study classroom tests and quizzes if you need more help or if you have tried to understand certain lectures in your IB chemistry course. If you want more help, read our article on the best IB Chemistry books to find additional study resources. You should learn the material during the school year and not cram just before the IB chemistry papers. Do you want to create the best possible application for university? We can help. PrepScholar Admissions is the world's best admissions advisory service. We combine the world's admissions advisors with our data-driven, proprietary admissions strategies. We've overseen thousands of students who have made it to their best schools, from state colleges to the Ivy League. We know what kinds of college students they want to admit. We want you accepted into your school of yours. Learn more about PrepScholar Admissions to maximize your chances of getting it. Common Mistakes IB Chemistry Students Make While Studying Many students struggle with IB Chemistry SL/HL. There are so many subjects you can learn and you can't lag behind. You need to learn during the school year to make ace IB chemistry papers. Common mistakes students make while studying are: #1: Avoid topics that you didn't fully understand in the classroom. If you haven't learned this in class, you need to seek further help, either through this IB chemistry study guide, IB chemistry books or through tutoring. #2: Only study a week or two before the IB chemistry exam. Exist too many topics to master in just a week or two (which is why the course is taught for one to two years). So, master the subjects as you learn them in class. Use this study guide if you need more help understanding the topics you cover in your class. Otherwise, you'll be just as nervous as this kid during the test. Core - 95 hours for SL and HL Both IB Chemistry SL and HL have the same basic requirements. They consist of 95 hours and cover 11 topics below. Topic 1: Stoichiometric Relationships - 13.5 Hours for SL and HL Notes on the Concept of Moles and Avogadro's Constant Notes on All Stoichiometric Videos and Notes 1.3: Reacting Matter Notes and Notes on Volumes Topic 2: Atomic Structure - 6 Hours for Videos and Notes on atomic structure SL and HL 2.2: Notes on electron configuration Topic 3: Periodicity - 6 hours for videos and notes on periodicity SL and HL 3.1: Periodic table of notes 3.2: Notes on periodic trends Topic 4: Chemical bonding and structure - 13.5 hours for SL and HL 4.1: Ionic bonding and structural notes 4.2: Covalent bonds 4.3: Covalent Structure Notes 4.4: Intermolecular Force Notes 4.5: Metal Bonding Notes Topic 5: Energy/Thermochemistry-9 Hours for SL and HL Energy Videos and Notes 5.1 : Measuring Energy Changes Note 6: Chemical Kinetics - 7 Hours for Kinetic Videos and Notes 6.1: Collision Theory and Reaction Rate Topic 7: Equil - 4.5 Hours for SL and HL Equilibrium Videos and Notes Topic 8: Acids and Principles-6.5 Hours for SL and HL Acids and Principles Study Guide to Acids and Principles Videos and Notes 8.1: Acid Theory and Principles Notes 8.1 2: Acid Properties and Policy Notes 8.4: Strong and Weak Acids and Principles Notes 8.5: Acid Deposition Notes Topic 9: Redox Processes-8 Hours for SL and HL Oxidation and Reduction Study Guide Redox Processes Videos and Notes 9 .1: Oxidation and Reduction Notes 9.2: Electrochemical Cell Notes Topic 10: Organic Chemistry-11 Hours for SL and HL Organic Chemistry Study Guide Organic Chemistry Videos and Notes 10.1 : Basics of Organic Chemistry Note 10.2: Functional Group Chemistry Notes Topic 11: Measurement and Processing of Data-10 Hours for SL and HL Spectroscopic Identification of Organic Compounds Study Guide Measurement Videos and Notes 11.1: Uncertainties and Errors in Measurement and Results of Note 11.2: Graphic Techniques Note 11.3: Spectroscopic Identification of Organic Compounds Notes one of the most important parts of your college application is what classes you choose to take in high school (in conjunction with how well you do in those classes). Our PrepScholar team of admissions experts have compiled their knowledge into this single guide for planning your high school course schedule. We will advise you on how to balance your schedule between regular and honors / AP / IB courses, how to choose extracurricular activities and what classes I can't afford not to take it. Another higher level (AHL)—60 hours for HL You will study only ten topics listed below if you are in IB Chemistry HL; standard level does not cover these topics. Topic 12: Atomic Structure - 2 Hours Atomic Structure Study 12.1: Electrons in Atoms Notes Topic 13: Periodic Table: Transition Metals - 4 Hours Periodic Table Study Guide 13.1: First Line d-block elements note 13.2: Color complexes notes Topic 14: Chemical bonding and structure – 7 hours Chemical bonding and structural notes 14.1: Covalent bonding and electron domain and molecular geometry Notes 14.2: Hybridization Notes Topic 15: Energy/Thermochemisthy-7 Hours 15.1: Energy Cycles Notes 15.2: Entropy and Spontane notes on the topic 16: Chemical Kinetics – 6 Hours Videos on Chemical Kinetics 16.1: Notes on expression and reaction mechanism speed 16.2: Activation energy notes Okay, you may not get to it. Topic 17: Equilibrium – 4 hours 17.1: Equilibrium Law Notes Topic 18: Acids and Principles - 10 hours 18.1: Lewis acids and principles of Note 18.2: Calculations involving acids and principles notes Topic 19: Redox processes - 6 hours Notes on oxidation and reduction 19.1: Electrochemical article 20: Organic Chemistry-12 Hours 20.1: Types of Organic Reactions Notes 20.2: Synthetic Route Notes Topic 21: Measurement and Analysis-2 Hours Videos on Data Measurement and Processing 21.1: Spectroscopic Identification of Organic Compounds Notes Option-15 Hours for SL and 25 Hours for HL Unfortunately, There Are No Free Online Study Guides for Options, but check out our article on IB chemistry books to find books, which examine the possibilities of topics. Overall IB Chemistry Reviews IB Chemistry HL 31 Common Errors: Richard Thornley, author of this video, has several other useful videos on IB Chemistry SL and HL available for free on YouTube IB Chemistry Web: This site goes through syllabus in-depth and explains key definitions and facts you need to know. What's next? Want more insight into what you will learn in Chemistry IB? Then take a look at our step-by-step guide to IB chemistry: SL and HL and our tips on balancing chemical equations. Prep book can be a very useful study tool. Read our guide to which are the best IB chemistry textbooks. How much do you know about the chemical properties of everyday things? Discover how to use muriatic acid to remove rust from the pot and pan and the effect of adding and removing certain ingredients to create the ultimate slime. You're hoping to squeeze into extra IB classes? Learn about the IB courses offered online. How far are you from 4.0? Use our simple GPA tool to determine how well you need to do in future classes to get gpa up to this magic number. An overview of all the basic topics for a higher level of study, written in an easy way to follow. Lots of practice including more than 100 Learning Check questions to make sure you understand key concepts, plus 40 additional exam-style questions to test your knowledge.Frequent error labels highlight mistakes students often make in exams. Classic question labels identifying popular types of questions in the IB exam. Exam Trap Labels indicating things to keep an eye on the exam that can trip you. Clear diagrams, charts, and tables to understand key concepts. Mole ConceptUnitsMole – Particle ConversionsMolar MassMole – Mass ConversionsBalancing EquationsMole Relationships in a Chemical ReactionMass Relationships in a Chemical ReactionLimiting ReactantDetermination of Formulae – Gravimetric AnalysisMixs & SolutionsMaking Solutions and Determination of Solute MassKinetic Molecular TheoryPressure Boyle's LawPressure LawCharles' LawConodál Combined Gas LawAvogadro is the LawAvogadro Law on Combining VolumesIdeal Gas LawYield: Theoretical, Experimental and percentage and molecular formulasCalculation from percentage informationCalculation from empirical dataSounds containing oxygenmolecular formulasSumotomy questionsSubatomy particlesNuclear atoms & isotopesIsotope notationsSatic spectrometersSatic atomic massCalculation relative mean atomic massCalculation Natural abundanceSusutomous Atom – Basic structureDiscovery for shells - Hydrogen spectrumExplaining hydrogen spectrumCommunical atoms - shells and subterroots - energy of light-emitters & Hund's RuleSignificant Filling Pattern in D-BlokSucesive Ionization Energy - Evidence for SubshellsSummary QuestionsStructure of the Periodic TableEffective Nuclear ChargeTrends in the Periodic Table – The Basic IdeaAtomic RadiusIonic RadiusAtomic vs. Ionic radiiionization energy – general trendElectric affinityElecterentegativityChemical propertiesReaction of alkaline metals and halogensRection of alkaline metals with waterAgainst halogensHalogens & HalidesChemical Properties Across Period 3 Oxide Period 3 Reactions with WaterFirst Line D-Block ElementsCatalystsComplex IonsSoul Compounds Transient MetalSummary QuestionsAdoration vs. StructureCete ruleTypes of bondingion hubs & FormulaeOne immonial compound - From the periodic tableTransition metal compoundis +2 or 2 +? Compounds of polyatomic ionsion compounds & propertiesMetallic bonding & alloymelting pointMerging point group 1 – Alkali MetalsCovalent BondingCovalent FormulaeLewis StructureVSEPR Theory & Molecular Shapes IEx Transit Lewis structureSums of molecules IIHybridizationTypes of resonance of covalent bonds and delokalizationSmyzenzace stabilization of energybond angle and double bondsformal chargeD polarity and molecular polarityAllotropy length and force ForcesLondon Disperzní SílyDipole - Dipole ForcesDipole - Indukované DipoleHydrogen Bonding Vrstvení intermolekulárních silJak si určit typ MMF? Typy solidsFyzikální vlastnosti solidsSummary QuestionsEndothermic & ExothermicPotenciální energetické diagramy (Enthalpy diagramy)Standardní podmínkyCalorimetrie – Entalalystatické výpočtySousesumptions v Calorimetry Calorimeter ConstantHess 'LawHeat of Hydration Bond EnthalpyStandard Enthalpy ZměnyStandardní Enthalpy formaciStandardní Enthalpy spalováníKteré vzorec mám použít? Born - Haber Cycles & Ionic Lattice EnthalpyEntropy Hess 'zákon a EntropySpontánnost & Gibb je zdarma EnergyFree Energy & Equilibrium Shrnutí OtázkySítrné reakční experimentyCollision TheoryFactors Affecting Rate of ReactionRate Expression/Rate LawThe Rate ConstantReaction MechanismsAktivace EnergySummary OtázkyDynamické rovnováhy přístup rovnovážnýPozice rovnováhy Koncentrace Rovnováha OtázkySummary OtázkySumatorizace kyselin a bázíků Acids & BasesH+ (AQ) nebo H3O+ (AQ)? Conjugate PairsAmphoteric SubstancesStrong & WeakDistinguishing Between Strong and WeakpH ScaleAcid DepositionCalculation of pH of Weak Acid SolutionsHydrolysis of SaltsBuffer SolutionsAcid – Base Titration – Strong Acid & Strong BaseAcid – Base Titration – Weak & StrongIndicatorspH CurvesRelative Strengths of Acids & BasesSummary Questions Oxidation NumbersRedox ReactionsReactivityHalf ReactionsDissolved Oxygen and BOD DefinitionsVoltaic CellsSalt BridgeStandard Electrode PotentialsElectrolysis of Molten SaltsElectrolysis of SolutionsFactors Affecting Amount of Electrolysis ProductBalancing Redox Reactions in Acidic SolutionSummary QuestionsHydrocarbonsNaming HydrocarbonsIsomersStructural IsomersBranching in Organic ChemistryAlkanesOrganic Functional GroupsFunctional Groups vs. Classes of CompoundsFunctional Group IsomerismAlkenes – Electrophilic Addition ReactionsElectrophilic Addition MechanismBenzene and DerivativesImportant IsomersHalogenoalkanesNucleophilic Substitution ReactionsSN1 – Nucleophilic Substitution – First Order KineticsSN2 – Nucleophilic Substitution – Second Order KineticsNucleophilic Substitution Reactions IIAlcoholsCombustion of AlcoholsOxidation of AlcoholsAlcohol Oxidation ProductsReduction of Aldehydes, Ketones and AcidsFunctional Groups Containing NitrogenElimination ReactionsEster Condensation ReactionsCondensation PolymersElectrophilic Substitution ReactionsReaction PathwaysStereoisomerismOptical IsomerismSummary QuestionsObservationsUncertainty in MeasurementsPrecision & AccuracySignificant DigitsRounding of NumbersAdding and Subtracting Significant DigitsMultiplying and Dividing Significant DigitsSystematic Uncertainty Absolute & NejistotaPropaguj c nejistotaTypy vztah'Formul compoundmass spectroscopy (MS)Jadern' magnetick' rezonanceInfrared SpectroscopyX-Ray Krystalography Krystalography