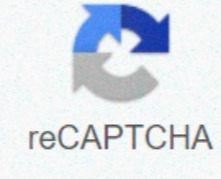


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Section 1 stability in bonding worksheet answers chapter 19

1 Chapter 19 Chemical Effects 2 Stability in Binding Some elements combine chemically and no longer have the same properties they did before forming a compound. A chemical formula consists of symbols and subscripts that indicate the number of atoms of an element in a compound. 3 Atoms form compounds when the compound is more stable than the separate atoms. Noble gases are more chemically stable than other elements because they have a complete outer energy level. Elements that do not have full outer energy levels are more stable in compounds. Atoms can lose electrons, get or share electrons to get a stable outer energy level. A chemical tape is the force that holds atoms together in a compound. 4 Types of Effects an ion is a charged particle because it has more or fewer electrons than protons. When an atom loses an electron, it becomes a positively charged ion; a superscript indicates the charge. When an atom gets an electron, it becomes a negatively charged ion. An ionic connection is held together by the ionic bond — the power of attraction between opposite charges of the ion. The result of this tape is a neutral compound. The sum of the charges on the earthhouses is zero. 5 Ionic Effects Ionic Effects Form When a Metal Is Added to a Non-Metal and Electrons Are Transferred Sodium+ Chlorine Magnesium + Oxygen Aluminum + Nitrogen Strontium + Fluorine 6 Molecules are neutral particles formed due to the share of electrons. A covalent bond is the power of attraction between atoms sharing electrons. Atoms can form double or triple effects depending on whether they share two or three pairs of electrons. Electrons shared in a molecule are kept more closely to the atoms with the larger core. A polar molecule has one side that is slightly negative and one end that is slightly positive, although the overall molecule is neutral. In a nonpolar molecule, electrons are shared equally. 7 Covalent Bonds Covalent tapes occur when a non-element ties to another non-linguistic electrons shared Boron + Nitrogen Carbon + Chlorine Hydrogen + Oxygen 8 Writing Chemical Formulas and Naming ConnectionChemists use symbols of the periodic table to write formulas for compounds. Binary connection- consists of two elements Use the name of the first element, the root name of the second element, and the suffix -ide to write the name of a binary ionic connection. Sodium Chloride Magnesium Oxide Lithium Fluoride Aluminum Nitride Calcium Sulfide 9 Oxidation Number—How Many Electrons Gained, Lost, or Shared Stable Oxidation Numbers and Their Least Common Multiple Formulas to Write. When writing formulas, remember that the compound is neutral. A formula should have the correct number of positive and negative onions so that the Balance. 10 Oxidation numbers and Dot Diagrams 11 Hydrate compound with water chemically attached to syions Polyatomic ion-positive or Loaded, covalently bound group atoms The connection contains more than one element. To write names, first write the name of the positive ion; then write the name of the negative ion. To write formulas, use the oxidation numbers, their least common multiple, and put brackets around the polyatomic ion before a subscript hydrate-connection to water chemically attached to its ion Name binary covalent compounds by using prefixes to indicate how many atoms of each element are in the compound. 12 Thanks for your participation! Transcription Section 1 Chemical Effects Stability in Binding A. Some elements combine chemically and no longer have the same properties they did before forming a compound. Chapter 19 Underlined words and phrases must be filled in by students on the Note-Taking Worksheet. B. A(n) chemical formula consists of symbols and subscripts indicating the number of atoms of an element in a compound. C. Atoms form compounds when the compound is more stable than the separate atoms. 1. Noble gases are more chemically stable than other elements because they have a complete outer energy level. 2. Elements that do not have full outer energy levels are more stable in compounds. 3. Atoms can lose, get or share electrons to get a stable outer energy level. 4. A(n) chemical tape is the force that holds atoms together in a compound. DISCUSSION QUESTION: Why do elements form connections? To become more chemically stable by getting a complete outer energy level Section 2 Types of Effects A. A(n) ion is a loaded state because it has more or fewer electrons than protons. 1. When an atom loses an electron, it becomes a positively charged ion; a superscript indicates the charge. 2. When an atom gets an electron, it becomes a negatively charged ion. B. An ionic connection is held together by the ionic bond — the power of attraction between opposite charges of the ion. 1. The result of this tape is a(n) neutral compound. 2. The sum of the charges on the onions is zero. C. Molecules are neutral particles formed due to the sharing of electrons. 1. A covalent bond is the power of attraction between atoms that share electrons. 2. Atoms can form double or triple effects depending on whether they share two or three pairs of electrons. T2 Chemical Effects Copyright © Glencoe/McGraw-Hill, a division of the McGraw-Hill Companies, Inc. Teacher Support & Planning Content Breakdown for Teaching 3. Electrons shared in a molecule are kept more closely to the atoms with the larger core. 4. A(n) polar molecule has one side that is slightly negative and one side that is slightly positive, although the overall molecule is neutral. 5. In a(n) nonpolar electrons are shared equally. BOOKING QUESTION: What two types of atomic bands form and do they differ? Ionic effects form when atoms lose or get electrons; covalent effects form when atoms share electrons. Part. 3 Writing formulas and Naming Connections A. Chemists use symbols from the periodic table to write formulas for connections. B. Binary connection-consists of two elements 1. Oxidation number - how many electrons an atom gained, lost or shared to become stable 2. Use oxidation numbers and their least common multiples to write formulas. A. When writing formulas, remember that the compound is neutral. B. A formula must have the correct number of positive and negative onions so that the charges balance. Copyright © Glencoe/McGraw-Hill, a division of the McGraw-Hill Companies, Inc. 3. Use the name of the first element, the root name of the second element, and the suffix -ide to write the name of a binary ionic connection. C. Polyatomic ion - positive or negatively charged, covalently bound group atoms 1. The connection contains three or more elements. 2. To write names, write the name of the positive ion first; then write the name of the negative ion. 3. To write formulas, use the oxidation numbers, their least common multiple, and put brackets around the polyatomic ion before adding a subscript. D. Hydrate-connection to water chemically attached to its ion E. Name binary covalent compounds by using prefixes to indicate how many atoms of each element are in the compound. Booking question: What is a binary connection? One that is formed from two elements Chemical Effects T3 Teacher Support & Planning Content Breakdown for Teaching (ongoing) (ongoing)

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